

Chem 115 POGIL Worksheet - Week #7 - Answers
First Law, Enthalpy, Calorimetry, and Hess's Law

Key Questions and Exercises

1. What are the two major ways in which the internal energy of an object can be categorized? How do these ways differ from one another?

An object can have kinetic energy and potential energy. The kinetic energy arises from motion of the object itself or the fundamental particles of which it is composed. The potential energy arises from its position and composition.

2. If a system loses heat, where does it go?

Because of the First Law of Thermodynamics, heat lost by the system is gained by its surroundings, and vice versa.

3. Describe the following processes as exothermic or endothermic:

- | | |
|---|-------------|
| a. an ice cube melts on a warm surface | endothermic |
| b. water freezes | exothermic |
| c. calcium chloride is mixed with water, resulting in a very hot solution | exothermic |
| d. ammonium nitrate is mixed in water, resulting in a very cold solution | endothermic |

4. The thermochemical equation for the burning of one mole of benzene under standard conditions is



- a. Is this reaction exothermic or endothermic? exothermic
- b. How much heat is released when a 5.00-g sample of benzene is burned in excess oxygen under standard conditions? (m.w. $\text{C}_6\text{H}_6 = 78.11 \text{ u}$)

$$q = (5.00 \text{ g}) \left(\frac{\text{mol}}{78.11 \text{ g}} \right) \left(\frac{-3267.7 \text{ kJ}}{\text{mol}} \right) = -209 \text{ kJ}$$

5. The specific heat of iron (at. wt. = 55.85 u) is 0.450 J/g·K. What is the heat capacity of a 23.5-g block of iron? What is the value of the molar heat capacity of iron?

$$C = (0.450 \text{ J/g}\cdot\text{K})(23.5 \text{ g}) = 10.5_{75} \text{ J/K} = 10.6 \text{ J/K}$$

$$C_m = (0.450 \text{ J/g}\cdot\text{K})(55.85 \text{ g/mol}) = 25.1_{325} \text{ J/mol}\cdot\text{K} = 25.1 \text{ J/mol}\cdot\text{K}$$

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DP Hallahan

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