

Naming Ionic Compounds

What are the structural units that make up ionic compounds and how are they named?

where

When working in chemistry, it is often convenient to write a chemical in symbols. For example we might write down the substance table salt as NaCl. In talking about chemistry however, it is a bit tacky to say "en-ay see-ell" when we want to refer to a substance. Also, in formal writing we should use the name of the compound rather than its symbols. Therefore we need to learn how to say the proper names of ionic substances.

Model 1 – Ion Charges for Selected Elements

1	H ⁺																		
2	Li ⁺	Be ²⁺									N ³⁻	O ²⁻	F ⁻						
3	Na ⁺	Mg ²⁺	Transition elements						Al ³⁺		P ³⁻	S ²⁻	Cl ⁻						
4	K ⁺	Ca ²⁺	Fe ²⁺ Fe ³⁺	Ni ²⁺ Ni ³⁺	Cu ⁺ Cu ²⁺	Zn ²⁺								Br ⁻					
5	Rb ⁺	Sr ²⁺			Ag ¹⁺						Sn ²⁺ Sn ⁴⁺			I ⁻					
6		Ba ²⁺				Hg ₂ ²⁺ Hg ²⁺					Pb ²⁺ Pb ⁴⁺								
	Cations														Anions				

1. Based on the information in Model 1:

- Based on the information in Model 1:
- Identify three elements that form only one cation. $\leftarrow$ Positive Ion
 Na^+ , Sr^{2+} , Al^{3+} (Any box with one element works.)
 - Identify three elements that form only one anion. $\leftarrow$ Negative Ion
 P^{3-} , O^{2-} , Br^- (Any box with one element works.)
 - Identify three elements that form more than one cation.
 Fe^{+2} and Fe^{+3} , Ni^{+2} and Ni^{+3} , Cu^{+1} and Cu^{+2}
 - In what region of the periodic table are these "multiple ion" elements usually located?
 Transition elements or "d" block

2. Consider the ions of potassium (K) and sulfur (S). Write chemical formulas for all possible ionic compounds involving these ions, using the simplest ratio(s) of potassium (K) and sulfur (S). Keep in mind that the sum of the charges in an ionic compound must equal zero.
- $2K^+$ and S^{2-} or K_2S
3. Consider the ions of iron (Fe) and sulfur (S). Write chemical formulas for all possible ionic compounds involving these ions, using the simplest ratio(s) of iron (Fe) and sulfur (S). Keep in mind that the sum of the charges in an ionic compound must equal zero.



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Table of Contents Pogil Periodic Table Extension Questions Answer Key

1. Understanding the eBook Pogil Periodic Table Extension Questions Answer Key
 - The Rise of Digital Reading Pogil Periodic Table Extension Questions Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Periodic Table Extension Questions Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Pogil Periodic Table Extension Questions Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Periodic Table Extension Questions Answer Key
 - Personalized Recommendations
 - Pogil Periodic Table Extension Questions Answer Key User Reviews and Ratings
 - Pogil Periodic Table Extension Questions Answer Key and Bestseller Lists

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 - Pogil Periodic Table Extension Questions Answer Key Budget-Friendly Options
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 - Pogil Periodic Table Extension Questions Answer Key Enhanced eBook Features
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 - Highlighting and Note-Taking Pogil Periodic Table Extension Questions Answer Key
 - Interactive Elements Pogil Periodic Table Extension Questions Answer Key
8. Staying Engaged with Pogil Periodic Table Extension Questions Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pogil Periodic Table Extension Questions Answer Key
9. Balancing eBooks and Physical Books Pogil Periodic Table Extension Questions Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pogil Periodic Table Extension Questions Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Pogil Periodic Table Extension Questions Answer Key
 - Setting Reading Goals Pogil Periodic Table Extension Questions Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Pogil Periodic Table Extension Questions Answer Key
 - Fact-Checking eBook Content of Pogil Periodic Table Extension Questions Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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