

PHYSICAL KINETIC MOLECULAR THEORY

Learning Objectives

- to identify theoretical differences in how gases are represented in the kinetic theory and in real-world situations;
- to explain the relationship of the postulates of the kinetic molecular theory;
- to explain how temperature affects molecular motion;
- to represent particle behavior using a particle model and a graph of average molecular speed.

Model 1: Representation of Gases in Different Phases

The following figure illustrates how the representative particles in gases behave in different phases of matter.

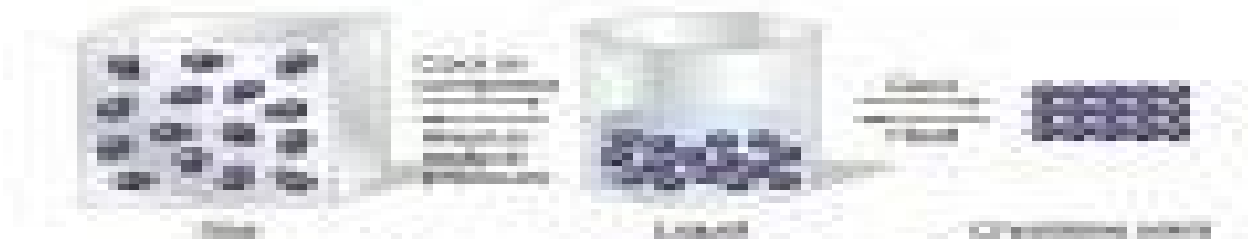


Figure 1: Representative particles in different phases of matter

Model 1 represents the behavior of particles in different phases of matter. The particles in the gas phase are moving rapidly in all directions, while the particles in the liquid phase are moving more slowly and in a more ordered fashion. The particles in the solid phase are moving very slowly and in a fixed position.

Key Questions

- What are the key differences in the behavior of particles in different phases of matter? How do these differences affect the properties of the substances? (See Model 1. The particles in the gas phase are moving rapidly in all directions, while the particles in the liquid phase are moving more slowly and in a more ordered fashion. The particles in the solid phase are moving very slowly and in a fixed position.)

Table 1: Differences in the behavior of particles

	PHASE 1	PHASE 2	PHASE 3
What are the key differences in the behavior of particles?	Particles in Phase 1 are moving rapidly in all directions, while the particles in Phase 2 are moving more slowly and in a more ordered fashion. The particles in Phase 3 are moving very slowly and in a fixed position.	Particles in Phase 2 are moving more slowly and in a more ordered fashion, while the particles in Phase 1 are moving rapidly in all directions. The particles in Phase 3 are moving very slowly and in a fixed position.	Particles in Phase 3 are moving very slowly and in a fixed position, while the particles in Phase 1 are moving rapidly in all directions. The particles in Phase 2 are moving more slowly and in a more ordered fashion.
How do these differences affect the properties of the substances?	These differences affect the properties of the substances in several ways. First, the particles in Phase 1 are moving rapidly in all directions, which means that they have a high average molecular speed. This results in a high pressure and a high temperature. Second, the particles in Phase 2 are moving more slowly and in a more ordered fashion, which means that they have a lower average molecular speed. This results in a lower pressure and a lower temperature. Third, the particles in Phase 3 are moving very slowly and in a fixed position, which means that they have a very low average molecular speed. This results in a very low pressure and a very low temperature.	These differences affect the properties of the substances in several ways. First, the particles in Phase 2 are moving more slowly and in a more ordered fashion, which means that they have a lower average molecular speed. This results in a lower pressure and a lower temperature. Second, the particles in Phase 1 are moving rapidly in all directions, which means that they have a high average molecular speed. This results in a high pressure and a high temperature. Third, the particles in Phase 3 are moving very slowly and in a fixed position, which means that they have a very low average molecular speed. This results in a very low pressure and a very low temperature.	These differences affect the properties of the substances in several ways. First, the particles in Phase 3 are moving very slowly and in a fixed position, which means that they have a very low average molecular speed. This results in a very low pressure and a very low temperature. Second, the particles in Phase 1 are moving rapidly in all directions, which means that they have a high average molecular speed. This results in a high pressure and a high temperature. Third, the particles in Phase 2 are moving more slowly and in a more ordered fashion, which means that they have a lower average molecular speed. This results in a lower pressure and a lower temperature.

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