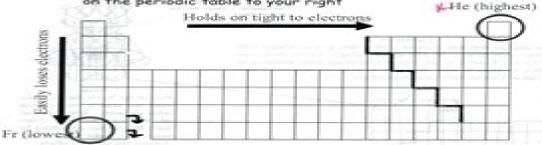
b. Why do elements get larger as you go down a group? Additional energy levels are added to the electron cloud as you go down a group. Land and no abnort

#### **Ionization Energy:**

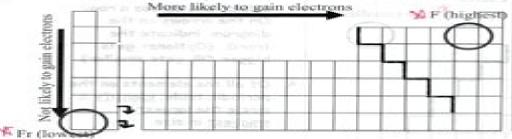
- 1. In simple terms, explain ionization energy. The energy required to remove an electron from a neutral atom.
- 2. Using the information in figure 5-15 (Pg 143) in your book. write the ionization energy for alkali metals & period #2 on the periodic table to your right



- On the arrows on the diagram. indicate the trend for ionization energy as you move down a column and across a row. (Options: easily loses an electron OR holds on tight to electrons)
- 4. Of all the elements on the periodic table, identify & circle the element with the highest & lowest ionization energy.
- Rank the following elements from highest to lowest ionization energy: a) Be, Ba, Mg = Be, Mg, Ba b) Li, F, B = F, B, Li
- 6. Why does it take less energy to remove an electron from a large atom compared to a small atom? Their electrons are closer to the nucleus, so the protons have a larger "pull" on them.

#### It takes less energy to remove an election from a large atom because Electronegativity: the effective nuclear charge decreases and electrons are

- 1. Explain Electronegativity: The ability of an atom to get electrons from other atoms.
- 2. Using the information in figure 5-20 (pg 151) in your book, write the electronegativity for the Halogens and period #2 on the periodic table below.



3. On the arrows on the diagram. fur Her indicate the trend for electronegativity as you move down a column and across a row. (Options: more likely to gain electrons, not likely to gain electrons)

WHILE BE

- 4. Of all the elements on the periodic table, identify & circle the element with the highest & lowest electronegativity.
- 5. Rank the following elements from most to least electronegative element: a) Be, Ba, Mg = Be, Mg, Ba b) Li, F, B - F, B, Li
- 6. Why are small atoms better at pulling electrons off of larger atoms than vice versa? Their electrons are closer to the nucleus, so the protons have a larger "pull" on the electrons in enough other atoms. The increased number of energy level(n) increases the distance over which nucleus must pull and therefore reduces the attraction for electrons

# **Pogil Cracking The Periodic Table Code Answer**

**Joshua BishopRoby** 

#### **Pogil Cracking The Periodic Table Code Answer:**

Atoms, Molecules & Elements: Patterns In the Periodic Table Gr. 5-8 George Graybill, 2015-10-01 This is the chapter slice Patterns In the Periodic Table from the full lesson plan Atoms Molecules Elements Young scientists will be thrilled to explore the invisible world of atoms molecules and elements Our resource provides ready to use information and activities for remedial students using simplified language and vocabulary Students will label each part of the atom learn what compounds are and explore the patterns in the periodic table of elements to find calcium Ca chlorine Cl and helium He through hands on activities These and more science concepts are presented in a way that makes them more accessible to students and easier to understand Written to grade and using simplified language and vocabulary and comprised of reading passages student activities crossword word search comprehension quiz and color mini posters our resource can be used effectively for test prep and your whole class All of our content is aligned to your State Standards and are written to Bloom's Taxonomy and Atoms, Molecules & Elements: The Periodic Table Gr. 5-8 George Graybill, 2015-10-01 This is the STEM initiatives chapter slice The Periodic Table from the full lesson plan Atoms Molecules Elements Young scientists will be thrilled to explore the invisible world of atoms molecules and elements Our resource provides ready to use information and activities for remedial students using simplified language and vocabulary Students will label each part of the atom learn what compounds are and explore the patterns in the periodic table of elements to find calcium Ca chlorine Cl and helium He through hands on activities These and more science concepts are presented in a way that makes them more accessible to students and easier to understand Written to grade and using simplified language and vocabulary and comprised of reading passages student activities crossword word search comprehension quiz and color mini posters our resource can be used effectively for test prep and your whole class All of our content is aligned to your State Standards and are written to Bloom's Taxonomy and STEM initiatives **Leveled Texts: The Periodic Table** Joshua BishopRoby, 2014-01-01 All students can learn about the periodic table through text written at four different reading levels Symbols on the pages represent reading level ranges to help differentiate instruction Provided comprehension questions complement the text Elements and the Periodic Table. Grades 5 - 8 Abbqy, 2013-01-02 Aligned to Common Core State Standards Elements and the Periodic Table present the basics of the Periodic Table in an easy to understand easy to master way It contains fun activities transparency masters quizzes tests rubrics grading sheets and more From basic elements to table organization Elements and the Periodic Table is the essential handbook for middle school science *Elements and the Periodic Table, Grades 5 - 12 Theodore S.* Abbqy, 2013-01-02 Aligned to Common Core State Standards Elements and the Periodic Table present the basics of the Periodic Table in an easy to understand easy to master way It contains fun activities transparency masters quizzes tests rubrics grading sheets and more From basic elements to table organization Elements and the Periodic Table is the essential handbook for middle school science **The Basics of the Periodic Table** Leon Gray, 2013-12-15 A sweeping history of

both the discovery and classification of elements and the development of the modern periodic table Included are discussions of the discovery of matter atoms atomic structure molecules compounds ions and isotopes as well as the first identifications of the 118 and counting elements and the various ways they have been classified and organized by prominent scientists up to the present day periodic table Instruction in how to read the periodic table is accompanied by examinations of the various groups of elements their location on the table and their properties and practical uses This text strongly supports Common Core Standards for the reading of scientific and technical texts and accounts and furnishes ample opportunities to summarize cite evidence and analyze connections between ideas individuals and events

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- 1. This book is structured into several chapters, namely:
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  - Chapter 2: Essential Elements of Pogil Cracking The Periodic Table Code Answer
  - Chapter 3: Pogil Cracking The Periodic Table Code Answer in Everyday Life
  - Chapter 4: Pogil Cracking The Periodic Table Code Answer in Specific Contexts
  - ∘ Chapter 5: Conclusion
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- 4. In chapter 3, this book will examine the practical applications of Pogil Cracking The Periodic Table Code Answer in daily life. The third chapter will showcase real-world examples of how Pogil Cracking The Periodic Table Code Answer can be effectively utilized in everyday scenarios.
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- 6. In chapter 5, this book will draw a conclusion about Pogil Cracking The Periodic Table Code Answer. The final chapter will summarize the key points that have been discussed throughout the book.

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