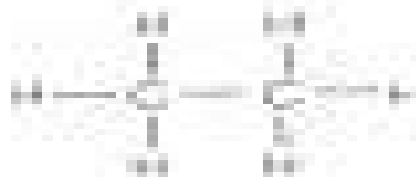


MICHEL 1: Relative electronegativities: $\text{O} > \text{N} > \text{C} > \text{H}$

A



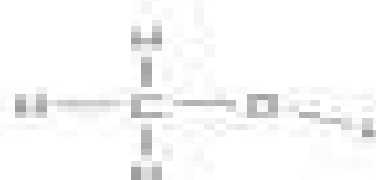
B



C



D



The bonds in B and C are polar.
 The molecules A and D are non-polar.

QUESTIONS

- For each molecule indicated 1, draw in the pairs of electrons shared between the atoms. Indicate which electrons are shared with hydrogen.
 - It may prove helpful distinguishing the atoms with dots that electrons shared in the molecule are shared between atoms.
 - If both atoms are equally electronegative, draw the electrons in the middle of the bond.
- For each molecule, draw the distribution of all bonding electrons that are shared unequally. Draw within a partial dipole moment in that sense, δ^+ or δ^- .
- Label all polar bonds with an arrow (↗).
- Label a polar/molecule as a polar molecule. Does the net dipole moment knowledge or a knowledge for electronegativity. Give the polarisation/molecule indicated and give the net dipole moment (↗).
- Class molecules that are polar/molecule as polar or non-polar molecules indicated? Be particularly well give an example.

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NARAYAN CHANGDER



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