

5. Examine the chemical formulas for the Arrhenius bases in Model 1. What anion do all the Arrhenius base chemical formulas have in common?

contain OH!

6. A student dissolved a small amount of baking soda in water and tested it with litmus paper. The litmus paper turned blue. Is baking soda likely an acid or a base?

BASE!

Read This!

In 1903 Svante Arrhenius won the Nobel Prize in Chemistry for defining acids and bases in terms of the ions produced. An Arrhenius acid is any substance that produces hydrogen ions [or hydronium ions (H_3O^+) a hydrogen ion attached to a water molecule] when dissolved in water. An Arrhenius base is any substance that produces hydroxide ions when dissolved in water. While the Arrhenius definitions of acids and bases is useful, it is limited. Johannes Brønsted and Thomas Lowry developed more general definitions for acids and bases using H^+ ion (proton) transfer as the focus.

Model 2 – Brønsted-Lowry Acids and Bases

Reaction 1	HCl(g) (acid)	+	$\text{H}_2\text{O(l)}$ (base)	$\rightleftharpoons$	$\text{H}_3\text{O}^+(\text{aq})$ + $\text{Cl}^-(\text{aq})$	gives accepts
Reaction 2	$\text{NH}_3(\text{aq})$ (base)	+	HF(aq) (acid)	$\rightleftharpoons$	$\text{NH}_4^+(\text{aq})$ + $\text{F}^-(\text{aq})$	gives accepts
Reaction 3	$\text{NH}_4^+(\text{aq})$ (acid)	+	$\text{H}_2\text{O(l)}$ (base)	$\rightleftharpoons$	$\text{NH}_3(\text{aq})$ + $\text{H}_3\text{O}^+(\text{aq})$	gives accepts
Reaction 4	$\text{F}^-(\text{aq})$ (base)	+	$\text{H}_2\text{PO}_4^-(\text{aq})$ (acid)	$\rightleftharpoons$	$\text{HPO}_4^{2-}(\text{aq})$ + HF(aq)	gives accepts

7. Identify the Brønsted-Lowry acids in Model 2.

a. Atoms of which element are present in all of the Brønsted-Lowry acids in Model 2?

H

b. How can you tell from Reaction 1 that HCl loses an H^+ ion rather than a hydrogen atom when the reaction occurs? *Hint: Look at the products.*

8. For each acid-base reaction in Model 2, describe the role of the Brønsted-Lowry acid in the H^+ ion (proton) transfer that occurs.

9. For each acid-base reaction in Model 2, describe the role of the Brønsted-Lowry base in the proton (H^+ ion) transfer that occurs.

Pogil Acids And Bases Answer Key

J Elliott



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Table of Contents Pogil Acids And Bases Answer Key

1. Understanding the eBook Pogil Acids And Bases Answer Key
 - The Rise of Digital Reading Pogil Acids And Bases Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Acids And Bases Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pogil Acids And Bases Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Acids And Bases Answer Key
 - Personalized Recommendations
 - Pogil Acids And Bases Answer Key User Reviews and Ratings
 - Pogil Acids And Bases Answer Key and Bestseller Lists
5. Accessing Pogil Acids And Bases Answer Key Free and Paid eBooks
 - Pogil Acids And Bases Answer Key Public Domain eBooks
 - Pogil Acids And Bases Answer Key eBook Subscription Services
 - Pogil Acids And Bases Answer Key Budget-Friendly Options
6. Navigating Pogil Acids And Bases Answer Key eBook Formats

- ePub, PDF, MOBI, and More
- Pogil Acids And Bases Answer Key Compatibility with Devices
- Pogil Acids And Bases Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pogil Acids And Bases Answer Key
 - Highlighting and Note-Taking Pogil Acids And Bases Answer Key
 - Interactive Elements Pogil Acids And Bases Answer Key
- 8. Staying Engaged with Pogil Acids And Bases Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pogil Acids And Bases Answer Key
- 9. Balancing eBooks and Physical Books Pogil Acids And Bases Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pogil Acids And Bases Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pogil Acids And Bases Answer Key
 - Setting Reading Goals Pogil Acids And Bases Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pogil Acids And Bases Answer Key
 - Fact-Checking eBook Content of Pogil Acids And Bases Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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