

Answers Pogil Activity: Glycolysis and the Krebs Cycle

1. Refer to Model 1.
 - a. What molecule from food is the primary reactant for glycolysis?
 - i. Glucose
 - b. How many carbon atoms are in that reactant molecule?
 - i. Six carbon atoms
2. The carbon atoms from glucose end up in pyruvate molecules as a product of glycolysis.
 - a. How many carbon atoms are in a pyruvate molecule?
 - i. Three carbon atoms
 - b. How many pyruvate molecules are made from each glucose molecule?
 - i. Two pyruvate molecules are made from each glucose molecule.
3. Does the process of glycolysis require an input of energy? Provide specific evidence from Model 1 to support your answer.
 - a. Glycolysis needs two ATP molecules to convert glucose.
4. Refer to Model 1. Propose an explanation for why the author of this activity put PGAL at the highest point in the Model 1 diagram.

The PGAL molecule has a higher potential energy than the glucose molecule. The energy from the ATP energy converting into ADP increase the potential
5. Does pyruvate have more or less potential energy than glucose? Provide specific evidence from Model 1 to support your answer.
 - a. Pyruvate has less potential energy than glucose, in the model we see an arrow that indicates the amount of potential energy in the step of the process. In the model, pyruvate is indicated to have a lower potential energy level.
6. What is the net production of ATP by glycolysis?
 - a. The net gain of ATP in glycolysis is two.
7. What molecule acts as an electron acceptor in glycolysis?
 - a. NADH

Pogil Glycolysis And The Krebs Cycle Answers

David Kirk



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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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 $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$
 one mole of HCl would be fully neutralized by one mole of NaOH

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web conclusion the identification of an unknown solid acid the purpose of this experiment was to determine the standardization of sodium hydroxide NaOH and oxalic acid $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ and the molar mass of an unknown solid acid using the molar mass of the standardized sodium hydroxide solution

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web conclusion in this lab we used titration to explore the concepts of stoichiometry and equivalence points we found the concentration of an unknown substance by mixing 2m hcl with the naoh of unknown concentration in order to experimentally ascertain the concentration of the naoh

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web this point of titration is referred to as the equivalence point considering the 1 1 stoichiometry of this acid base reaction NaOH aq $\text{C}_6\text{H}_4\text{COOH}$ aq $\text{C}_6\text{H}_4\text{COONa}$ aq H_2O l the point of equivalence is the point of titration when the number of moles of naoh na added is equal to the number of moles of khp nb in the