

## Mole Ratios

How can the coefficients in a chemical equation be interpreted?

### Why?

A balanced chemical equation can tell us the number of reactant and product particles (ions, atoms, molecules or formula units) that are necessary to conserve mass during a chemical reaction. Typically when we balance the chemical equation we think in terms of individual particles. However, in real life the reaction represented by an equation occurs an unimaginable number of times. Short of writing very large numbers ( $10^{23}$  or larger) in front of each chemical in the equation, how can we interpret chemical equations so that they more realistically represent what is happening in real life? In this activity you will explore the different ways a chemical reaction can be interpreted.

### Model 1 – A Chemical Reaction



1. Consider the reaction in Model 1.
  - a. What are the coefficients for each of the following substances in the reaction?



- b. Draw particle models below to illustrate the reaction in Model 1.



2. Consider each situation below as it relates to the reaction in Model 1.
  - a. Calculate the amount of reactants consumed and products made.
  - b. Record the ratio of  $\text{N}_2$  to  $\text{H}_2$  to  $\text{NH}_3$ . Reduce the ratio to the lowest whole numbers possible.

|                                                                                               | $\text{N}_2$<br>Consumed | $\text{H}_2$<br>Consumed | $\text{NH}_3$<br>Produced | Ratio $\text{N}_2:\text{H}_2:\text{NH}_3$<br>(reduced) |
|-----------------------------------------------------------------------------------------------|--------------------------|--------------------------|---------------------------|--------------------------------------------------------|
| For a single reaction, how many molecules of each substance would be consumed or produced?    | 1 molecule               | 3 molecules              | 2 molecules               | 1:3:2                                                  |
| If the reaction occurred one hundred times, how many molecules would be consumed or produced? | 100 molecules            | 300 molecules            | 200 molecules             | 1:3:2                                                  |
| If the reaction occurred 538 times, how many molecules would be consumed or produced?         | 538 molecules            | 1614 molecules           | 1076 molecules            | 1:3:2                                                  |

# Pogil Chemistry Mole Ratios Answers

**Lifelique**



## **Pogil Chemistry Mole Ratios Answers:**

*Mole Ratios* Lifeliqe,2019 This lesson plan relates balanced chemical equations to everyday analogies defines stoichiometry and uses mole ratios to convert between amounts of substances in chemical reactions **Counting Moles** Nigel P. Freestone,2016-01-04 Students studying chemistry often struggle with the mole Counting Moles provides an effective aid to learning by giving clear and confident presentation of the essentials of the mole concept needed by those starting chemistry courses This user friendly self teach e book is split into six chapters which sequentially introduce the mole calculating frame to help solve problems Over 200 fully worked examples are given along with several hundred questions The mole concept is applied to topics such as relative atomic mass and relative formula mass percentage composition empirical and molecular formula The book also covers concentration its units volumetric analysis and the relationship between volume mass and moles of gases Counting Moles culminates in you taking a Mole Driving Test On passing this test you are issued with a Counting Moles Driving License that will give you all the confidence required to correctly answer all mole calculations

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