

Experiment: Solubility Rules and Net Ionic Equations

This is to be written up in its entirety in your laboratory notebook

OBJECTIVES

In this experiment you will mix small quantities of aqueous ionic solutions together and in some cases, observe the formation of precipitates.

OVERVIEW

1. Mix small quantities of different ionic solutions together.
2. Observe the absence or presence of a precipitate and its physical characteristics such as color, amount, etc.
3. Formulate general solubility rules based on experimental observations.
4. Write net ionic equations to describe the precipitation reactions.

SAFETY

- Wear safety goggles at all times.
- Wear shoes that cover the entire foot.
- Never touch, taste or smell chemicals.
- Keep your hands away from your face, eyes, nose and mouth.
- Always wash your hands before leaving the laboratory.

PRELAB QUESTIONS

References: Ebbing, "General Chemistry", 5th ed. pp. 91-94 (Molecular-Ionic Equations), pp. 95-98 (Precip. Reactions)

1. Write the balanced equation for dissolving the following solid ionic salts in water:



2. Write the complete ionic and the balanced net ionic equations for the following reactions:



3. Use the general solubility rules to predict which of the following compounds are soluble in water and which are not.

a. NiCl_2

b. NaOH

c. Ni(OH)_2

d. NaCl

e. $(\text{NH}_4)_2\text{S}$

f. $\text{Pb(NO}_3)_2$

g. PbS

h. NH_4NO_3

i. CaSO_4

j. K_2CO_3

k. K_2SO_4

l. CaCO_3

m. KBr

n. FeCl_3

o. FeBr_3

p. KCl

Pogil Answer Key Solubility Rules And Net Ionic Equations

G Psacharopoulos



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