

This is like a distribution in math. $2(x+3) \rightarrow 2x+6$
 Everything inside of the parentheses is multiplied by two.

- b. What does the subscripted "2" outside the parentheses of the chemical formula tell you about the compound?

The subscript "2" outside of $\text{Ba}(\text{NO}_3)_2$ the parentheses indicate that there are two of the polyatomic ion nitrate (NO_3^-).

20. How many atoms of each element are in one formula unit of ammonium phosphate, $(\text{NH}_4)_3\text{PO}_4$?

nitrogen

3

hydrogen

12

phosphorus

1

oxygen

4

21. A student writes the chemical formula for the ionic compound calcium hydroxide as CaOH_2 .

- a. Write the chemical formula for each ion in the compound.

Calcium: Ca^{+2}

Hydroxide: OH^{-1}

- b. Why is the student's chemical formula for the compound calcium hydroxide wrong?

The subscript of "2" after hydrogen (H) indicate two hydrogens, not two hydroxide polyatomic ions. Parentheses should be placed around OH^{-1} and a subscript of "2" should be placed outside of the parentheses.

22. Many of the chemical formulas in Model 3 include parentheses. Which one of the following rules summarizes the appropriate use of parentheses in ternary ionic compounds? For the three rules that do not apply in all cases, show at least one counter example from the chemical formulas in Model 3.

Parentheses are used around any ion that is used more than once in a formula unit.

↳ Not true - $\text{Al}_2(\text{CO}_3)_3$ - Al has more than one atom.

Parentheses are used around any polyatomic ion.

↳ Not true - KNO_3 - NO_3^{-1} has no parentheses because there is only one.

Parentheses are used around any polyatomic ion used more than once in a formula unit.

Parentheses are only used around polyatomic anions used more than once in a formula unit.

↳ Not true - $(\text{NH}_4)_3\text{PO}_4$ - ammonium is a polyatomic cation. (NH_4^{+1})

23. Write chemical formulas for the following ternary ionic compounds.

- a. Calcium sulfate



- b. Copper(II) nitrate



- c. Lithium phosphate



- d. Potassium permanganate



- e. Aluminum sulfite



- f. Magnesium bicarbonate



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John Sadler, Mike Taylor



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Table of Contents Pogil Activities For High School Chemistry Solutions Answer Key

1. Understanding the eBook Pogil Activities For High School Chemistry Solutions Answer Key
 - The Rise of Digital Reading Pogil Activities For High School Chemistry Solutions Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Activities For High School Chemistry Solutions Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pogil Activities For High School Chemistry Solutions Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Activities For High School Chemistry Solutions Answer Key

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- Pogil Activities For High School Chemistry Solutions Answer Key User Reviews and Ratings
- Pogil Activities For High School Chemistry Solutions Answer Key and Bestseller Lists
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 - Pogil Activities For High School Chemistry Solutions Answer Key eBook Subscription Services
 - Pogil Activities For High School Chemistry Solutions Answer Key Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
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 - Pogil Activities For High School Chemistry Solutions Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pogil Activities For High School Chemistry Solutions Answer Key
 - Highlighting and Note-Taking Pogil Activities For High School Chemistry Solutions Answer Key
 - Interactive Elements Pogil Activities For High School Chemistry Solutions Answer Key
- 8. Staying Engaged with Pogil Activities For High School Chemistry Solutions Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pogil Activities For High School Chemistry Solutions Answer Key
- 9. Balancing eBooks and Physical Books Pogil Activities For High School Chemistry Solutions Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pogil Activities For High School Chemistry Solutions Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pogil Activities For High School Chemistry Solutions Answer Key
 - Setting Reading Goals Pogil Activities For High School Chemistry Solutions Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pogil Activities For High School Chemistry Solutions Answer Key

- Fact-Checking eBook Content of Pogil Activities For High School Chemistry Solutions Answer Key
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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