

Chapter 9

9.1 Compounds and Chemical Change

- **Atom**- the smallest unit of an element that can exist alone or in combination with other elements.
- **Compounds**- formed when atoms are held together by an attractive force called a **chemical bond**
- **Molecule**- The smallest particle of a compound or a gaseous element, that can exist and still remain the characteristic chemical properties of a substance
- **Chemical reaction**- a change in matter where **different chemical substances are created by breaking and/or forming chemical bonds.**
 - Examples: Photosynthesis
- Energy is *absorbed* -> produce new chemical substances -> more internal potential energy
- Energy is *released* when new chemical substances are produced -> less internal potential energy
- **Chemical Equation**: a way of describing what happens during a chemical reaction
 - Example: Photosynthesis equation-
Energy/Sunlight + carbon dioxide molecules + water molecules -> plant material molecules + oxygen molecules

9.2 Valence Electrons and Ions

- the number of electrons in the **outermost orbital** that **usually determines the chemical properties of an atom**
- **Valence electron**:- the number of electrons in the outermost orbital that usually determines the properties of an atom
- **Octet Rule**- after the element helium, atoms attempt to acquire an outer orbital with eight electrons through chemical reactions
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9.3 Chemical Bonds

- **Chemical bond**-an attractive force that holds atoms together in a compound.
- Three Classes of Chemical Bonds-
 - Ionic Bonds
 - Covalent Bonds
 - Metallic Bonds
- **Ionic Bonds**- type of linkage formed from the electrostatic attraction between oppositely charged ions in a chemical compound
- **Covalent bonds**- result when atoms achieve the noble gas electron structure by *sharing* electrons. They generally formed between the nonmetallic elements on the right side of the periodic table
- **Metallic bonds**- formed in solid metals such as iron, copper, and the other metallic elements that make up about 80 percent of all the elements.
- **Heat Formation**- energy that is released during an ionic bond formation
- **Formula**- describes what elements are in the compound and in what proportions
- **Single bond**- covalent bond
- **Double bond**- formed by two pairs of electrons that are shared by two atoms

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