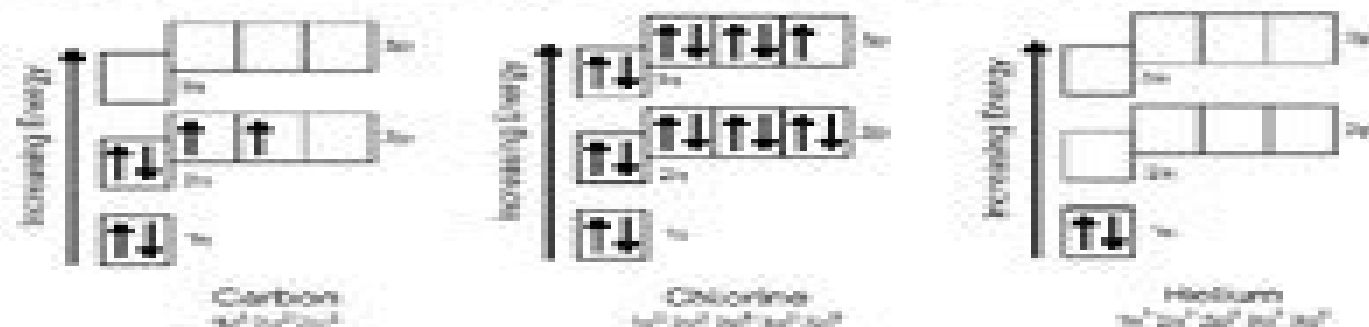


Model II: Electron Configuration with Orbital Diagrams



162 In orbital diagrams, what symbols do you think indicate the following?

- Electrons $\rightarrow$ arrows
- Orbitals $\rightarrow$ boxes
- Sub-level $\rightarrow$ boxes that connect

163 Based on the diagram above, how many electrons does chlorine have? How can you tell?

17, number of arrows

164 When will an electron enter the $2p$ orbital?

After the $2s$ sublevel is filled

165 How do orbital diagrams show the Aufbau principle?

No sublevel receives electrons unless the ones below it are already full.

166 Speculate with your group about why the electrons in carbon's $2p$ sub-level might arrange themselves as depicted above?

e^- have $\ominus$ charge & would rather spread apart if possible

Pogil Hybrid Orbitals

Carl Johan Ballhausen, Harry B. Gray



Pogil Hybrid Orbitals:

Orbitals in Chemistry Victor M. S. Gil, 2000-08-10 This text presents a unified and up to date discussion of the role of atomic and molecular orbitals in chemistry from the quantum mechanical foundations to the recent developments and applications The discussion is mainly qualitative largely based on symmetry arguments It is felt that a sound mastering of the concepts and qualitative interpretations is needed especially when students are becoming more and more familiar with numerical calculations based on atomic and molecular orbitals The text is mathematically less demanding than most traditional quantum chemistry books but still retains clarity and rigour The physical insight is maximized and abundant illustrations are used The relationships between the more formal quantum mechanical formalisms and the traditional chemical descriptions of chemical bonding are critically established This book is of primary interest to undergraduate chemistry students and others taking courses of which chemistry is a significant part *Atomic and Molecular Orbitals* George I. Sackheim, 1966 Orbitals: With Applications In Atomic Spectra Charles Stuart Mccaw, 2015-03-30 This book describes atomic orbitals at a level suitable for undergraduates in chemistry The mathematical treatment is brought to life by many illustrations rendered from mathematical functions no artists impressions including three dimensional plots of angular functions showing orbital phase and contour plots of the wavefunctions that result from orbital hybridisation Orbitals extends the key fundamental quantum properties to many electron atoms linear combinations of atomic orbitals simple molecules delocalised systems and atomic spectroscopy By focusing on simple model systems use of analogies and avoiding group theory the results are obtained from initial postulates without the need for sophisticated mathematics **Atomic Orbitals** Yue-Ling Wong, 1997-03 This Captivating Interactive CD ROM Aids Students In Visualizing The Abstract Concept Of Atomic Orbitals Atomic Orbitals Will Effectively Assist Students Comprehension Of The Challenging Concepts Of Orbital Shape Hybrid Orbitals And The Filling Of Orbitals Through Its Scientific Visualization Techniques And 3D Computer Animations As A Self Paced Individualized Program Or As A Classroom Lecture Tool Atomic Orbitals Is Ideal For Use In All Course Levels From General Chemistry Through Physical Chemistry Orbital Interactions in Chemistry Thomas A. Albright, Jeremy K. Burdett, Myung-Hwan Whangbo, 1985-02 This advanced text on applied molecular orbital theory covers organic organometallic inorganic and solid state chemistry to demonstrate how common orbital situations arise throughout the whole chemical spectrum Highlights the relationships among orbitals enabling readers to see the theoretical connections across the traditional boundaries of chemistry subdisciplines thus simplifying structure and reactivity problems Hybrid Atomic Orbitals in Molecular Orbital Methods Martin Grayson, 1984 **The Importance of Antibonding Orbitals** Milton Orchin, Hans H. Jaffé, 1967 **Molecular Orbital Theory** C. J. Ballhausen, Carl Johan Ballhausen, Harry B. Gray, 1965 Pictorial Orbital Theory John Michael Tedder, Antony Nechvatal, 1985 The Organic Chemist's Book of Orbitals William Jorgensen, 2012-12-02 The Organic Chemist s Book of Orbitals focuses on the mechanisms stereochemistry and reactivity of

molecular orbitals Composed of four chapters the book outlines how molecular orbitals are created by delocalization Concerns include CC and CH single bond orbitals bond orbitals and group orbitals and the localized orbitals of CH₂ and CH₃ groups Schematic diagrams are presented to show the nature reactions and compositions of molecular orbitals The text offers a list of molecules and orbital occupancies Orbital drawings are presented to show the differences of the molecular orbitals of hydrogen water ammonia methane nitrogen carbon monoxide and acetylene The book also provides an index of references for the molecular geometries and orbital energies employed in the orbital drawings Considering the weight of data presented the book is a great find for readers interested in studying molecular orbitals *Orbitals and Symmetry* David Selway Urch,1970 **Supplement for Symmetry, Orbitals and Spectra** Milton Orchin,Hans H. Jaffé,1971 *A Survey of Bonding with Hybridized Orbitals* Virginia Lewis,1962 **Molecular Orbital Theory** Carl Johan Ballhausen,Harry B. Gray,1964 **Molecular Orbital and Natural Bond Orbital Studies of Some Transition Metal Polysulfides** David Robert Kanis,1988 *Orbitals and Chemical Bonding* Paul Francis Lynch,1966 **Principles of Atomic Orbitals** Norman Neill Greenwood,1968 **Frontier Orbitals and Reaction Paths** Kenichi Fukui,Hiroshi Fujimoto,1997-01-01 This book is a collection of selected papers on the Frontier Orbital Theory by Nobel prizewinner Kenichi Fukui Chemistry 1981 with introductory notes It provides the basic concept and formulation of the theory and the physical and chemical significance of the frontier orbital interactions in chemistry together with many practical applications The formulation of the Intrinsic Reaction Coordinate and applications to some simple systems are also presented The aim of this volume is to show by what forces chemical reactions are driven and to demonstrate how the regio and stereo selectivities are determined in chemical reactions Students and senior investigators will gain insight into the nature of chemical reactions and find out how quantum chemical calculations are connected with chemical intuition **Hybrid Atomic Orbitals in Valence Theory** Ernest Frederick Kirkwood,1977

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