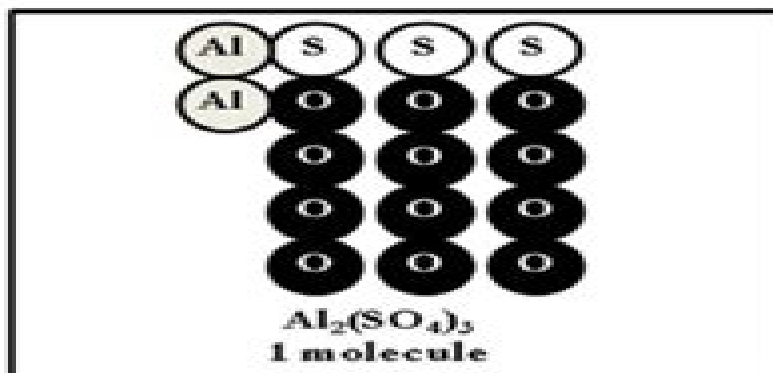


## Balancing Chemical Equations

### Why?

In a chemical reaction, matter cannot be created or destroyed. The mass of the reactants in a chemical reaction must be equal to the mass of the products. It is not possible to start a chemical reaction with 2 atoms of hydrogen and produce 2 atoms of gold. Similarly, it is not possible to start a chemical reaction with 2 atoms of hydrogen and produce only 1 atom of hydrogen. When a chemical equation is written, the same number of each type of atom must appear on both the product and reactant side.

### Model 1



- 1) According to Model 1, how many atoms of hydrogen and oxygen are in 1 molecule of water,  $\text{H}_2\text{O}$ ?
- 2) According to Model 1, how many atoms of hydrogen and oxygen are in 1 molecule of  $\text{H}_2\text{O}_2$  hydrogen peroxide?
- 3) According to Model 1, how many atoms of aluminum, sulfur, and oxygen are in 1 molecule of aluminum sulfate,  $\text{Al}_2(\text{SO}_4)_3$ ?

# Pogil Balancing Chemical Equations

**Rumi Michael Leigh**



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