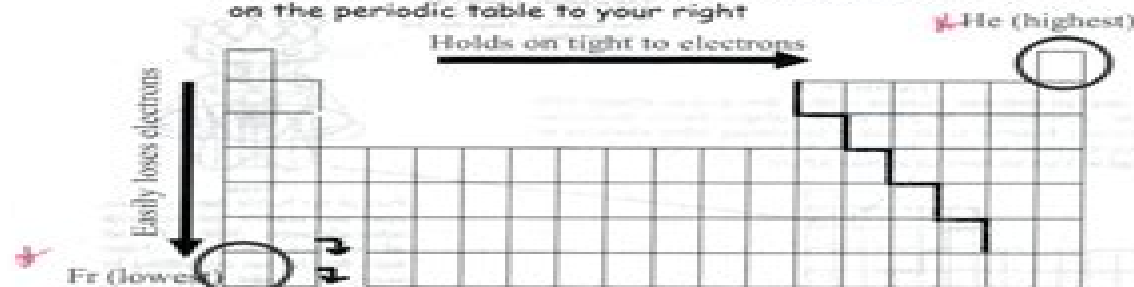


b. Why do elements get larger as you go down a group? Additional energy levels are added to the electron cloud as you go down a group.

Ionization Energy:

1. In simple terms, explain ionization energy. The energy required to remove an electron from a neutral atom.
2. Using the information in figure 5-15 (Pg 143) in your book, write the ionization energy for alkali metals & period #2 on the periodic table to your right



3. On the arrows on the diagram, indicate the trend for ionization energy as you move down a column and across a row. (Options: easily loses an electron OR holds on tight to electrons)

4. Of all the elements on the periodic table, identify & circle the element with the highest & lowest ionization energy.

5. Rank the following elements from highest to lowest ionization energy:

a) Be, Ba, Mg = Be, Mg, Ba

b) Li, F, B = F, B, Li

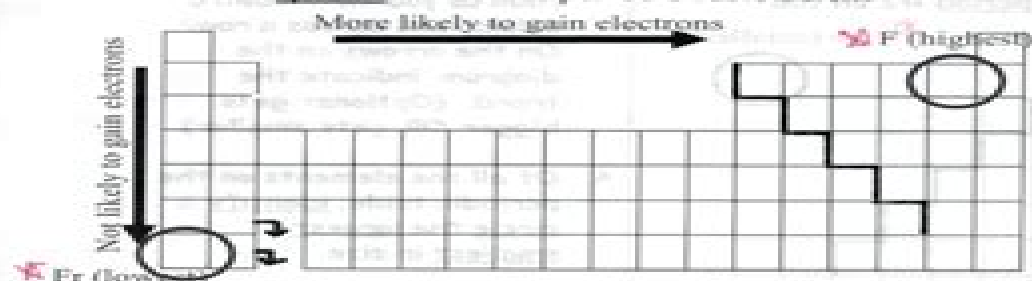
6. Why does it take less energy to remove an electron from a large atom compared to a small atom? Their electrons are closer to the nucleus, so the protons have a larger "pull" on them.

It takes less energy to remove an electron from a large atom because

Electronegativity: the effective nuclear charge decreases and electrons are not held as tightly since

1. Explain Electronegativity: The ability of an atom to get electrons from other atoms.

2. Using the information in figure 5-20 (pg 151) in your book, write the electronegativity for the Halogens and period #2 on the periodic table below.



3. On the arrows on the diagram, indicate the trend for electronegativity as you move down a column and across a row. (Options: more likely to gain electrons, not likely to gain electrons)

4. Of all the elements on the periodic table, identify & circle the element with the highest & lowest electronegativity.

5. Rank the following elements from most to least electronegative element:

a) Be, Ba, Mg = Be, Mg, Ba

b) Li, F, B = F, B, Li

6. Why are small atoms better at pulling electrons off of larger atoms than vice versa?

Their electrons are closer to the nucleus, so the protons have a larger "pull" on the electrons in other atoms.

not detailed enough
The increased number of energy level(s) increases the distance over which nucleus must pull and therefore reduces the attraction for electrons

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