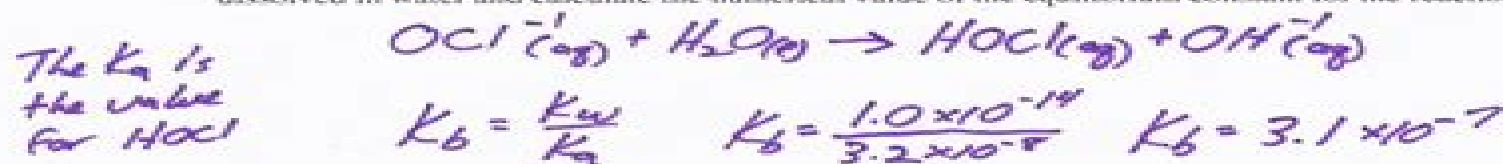


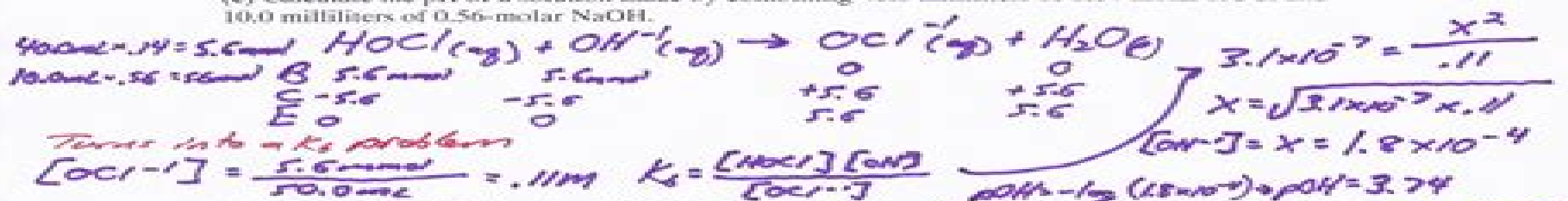
Key (5)

5) Continued

(b) Write the correctly balanced net ionic equation for the reaction that occurs when NaOCl is dissolved in water and calculate the numerical value of the equilibrium constant for the reaction.



(c) Calculate the pH of a solution made by combining 40.0 milliliters of 0.14-molar HOCl and 10.0 milliliters of 0.56-molar NaOH.



(d) How many millimoles of solid NaOH must be added to 50.0 milliliters of 0.20-molar HOCl to obtain a buffer solution that has a pH of 7.49? Assume that the addition of the solid NaOH results in a negligible change in volume.



$$\frac{K_a}{[\text{H}^+]} = \frac{\text{mmol OCl}^-}{\text{mmol HOCl}} \rightarrow \frac{3.2 \times 10^{-8}}{3.2 \times 10^{-8}} = \frac{x}{10. \text{ mmol} - x} \rightarrow x = 5 \text{ mmol OCl}^-$$

mmol HOCl = 50.0 mL $\times$.20 M = 10. mmol To create a 1 to 1 ratio, 5.0 mmol NaOH to convert to 5.0 mmol OCl⁻ (5.0 mmol HOCl)

(e) Household bleach is made by dissolving chlorine gas in water, as represented below.



Calculate the pH of such a solution if the concentration of HOCl in the solution is 0.065 molar.

* Trick question. If $[\text{HOCl}] = .065 \text{ M}$, then $[\text{H}^+]$ is also .065 M.

$$\text{pH} = -\log(.065)$$

$$\text{pH} = 1.2$$

* The hydrogen ions from HOCl will not influence the pH of the solution due to the presence of the strong acid HCl.

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Iris Stovall



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