

POGIL: Periodic Trends in Atomic Properties

Information

Properties such as the size of an atom (atomic radius), the energy required to remove an electron from an atom (ionization energy), and the energy associated with the addition of an electron to a gaseous atom (electron affinity) can be understood in terms of the electron configuration of the atom and a competition between *electron-nucleus attraction* and *electron-electron repulsion*.

Preliminary Activity - Complete the orbital diagrams for the ground state configuration.

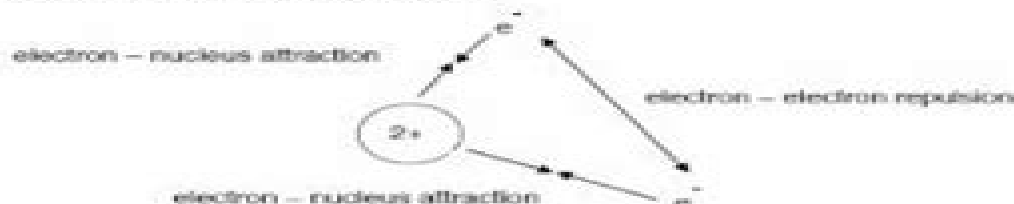
	1s	2s	2p		1s	2s	2p
H				C			
He				N			
Li				O			
Be				F			
B				Ne			

Model 1: Electron Ionization & Capture and Electrostatic Interactions

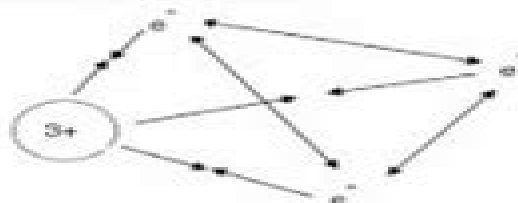
Reaction Equation for Electron Ionization: $X \rightarrow X^+ + e^-$ $\Delta E = 1^{\text{st}}$ ionization energy

Reaction Equation for Electron Capture: $X + e^- \rightarrow X^-$ $\Delta E = \text{electron affinity}$

Dominant inter-particle forces in helium:



Dominant inter-particle forces in lithium:



Pogil Periodic Trends In Atomic Properties

TD Snyder



Pogil Periodic Trends In Atomic Properties:

The Amazing World of the Atom and the Ultimate Periodic Table of the Elements Costel Rizescu, Mihaela Rizescu, 2018-05-24 This book is addressed to undergraduate and graduate students taking physics and chemistry courses at an introductory level STEM high school students and to more experienced professors and researchers who need to stay up to date with the most recent advances in data as well. It is also beneficial to those involved in materials science, medical physics, biotechnology, astronomy, quantum chemistry, Earth and space science and other related fields. The volume consists of two parts: the first part is devoted to the quantum mechanics theory of atomic structures as discovered and verified by physics, whereas the second part presents and analyzes a large amount of the latest critically evaluated data for the properties of all elements on the periodic table of the elements. Periodic trends for ionization energy, electronegativity, atomic radius, density, and melting points are graphed and thoughtfully analyzed. Additional data of ionic radius, covalent radius, van der Waals radius, and 12 coordination atomic radius for almost all elements on the periodic table are also presented and evaluated. The authors discussed over 30 problems, review questions, and their solutions to help deepen the insight in this subject area to make this book a real study text. Plenty of problems are given to elucidate the material. This book represents the culmination of the authors' many years of research and teaching physics, materials science, and instrumentation. It is also a crystallization of their passion and strong interest in the history of atomic physics and the philosophy of science in general. From the beginning, it was the authors' intention to include information in this book that is not easily located elsewhere or is not found completely in a single book or publication. Thus, while this book can be used as a text, the authors hope that it will be used as a useful reference too. Each chapter and subchapter of this book capitalizes on the strengths, comments, feedback, and criticism that the authors expect to have from students, faculty, and working professionals.

Atomic Properties of the Elements, 2001. *The periodic table. 4th ed.*, Atomic Theories F. H. Loring, 1921. **New Trends in Atomic Physics** Gilbert Grynberg, Raymond Stora, 1984. The Periodic System (1920 - 1923) Niels Bohr, 1977-01-15. Whilst volume 3 dealt primarily with the development of the quantum theory, The Periodic System contains in 2 parts the papers treating the application of the theory to the problem of atomic structure and the explanation of the periodic system of the chemical elements.

Development of a Periodic Function for Atomic Radius Jaswinder Kaur, Jaspreet Kaur, 2011-12. A number of experimental methods had been developed till date to evaluate and study the periodic properties. To study these periodic properties theoretically, we have developed a new linear approach. A mathematical periodic function has been developed to calculate atomic radii, covalent and van der Waals of any element by knowing just its atomic number. This function consists of a common equation along with additional terms, correction factor for different blocks s, p, d, and f of periodic table. This function has been derived mathematically by relating different factors in form of correction factors which affect the atomic radii using graphs. It takes into account the chemical behaviour of all the elements in their respective groups and periods.

Atoms

and Their Spectroscopic Properties Viacheslav Shevelko, 2014-01-15 **Atomic Structure and Periodicity** Jack Barrett, 2002 Each text in this series provides a concise account of the basic principles underlying a given subject embodying an independent learning philosophy and including worked examples This text covers atomic structure and periodicity

Periodic Trends Governing the Interactions Between Impurity Atoms [H-Ar] and [alpha]-U., 2008 The binding energies geometries charges and electronic structures of a series of impurity atoms H Ar interacting with the alpha U lattice in various configurations were assessed by means of density functional theory calculations Periodic trends governing the binding energy were highlighted and related to the electronic properties of the impurity atoms with some consideration given to the band structure of alpha U The strongest bound impurity atoms include C N O and Si P S The general trends in the binding energy can be reproduced by a simple parameterisation in terms of the electronegativity charge transfer and covalent radius elasticity theory of the impurity atom The strongest bound atoms deviate from this model due to their ability to bind with an optimum mixture of covalency and ionicity This last point is evidenced by the partial overlap of the impurity atom p band with the hybrid d f band of alpha U It is expected that the trends and general behaviour reported in this work can be extended to the interactions of impurity atoms with other metallic systems *Primer of the Atoms* Henry David Hubbard, 1937

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