

## Naming Molecular Compounds

How are the chemical formula and name of a molecular compound related?

### Why?

When you began chemistry class this year, you probably already knew that the chemical formula for carbon dioxide was  $\text{CO}_2$ . Today you will find out why  $\text{CO}_2$  is named that way. Naming chemical compounds correctly is of paramount importance. The slight difference between the names carbon monoxide ( $\text{CO}$ , a poisonous, deadly gas) and carbon dioxide ( $\text{CO}_2$ , a greenhouse gas that we exhale when we breathe out) can be the difference between life and death! In this activity you will learn the naming system for molecular compounds.

### Model 1 – Molecular Compounds

| Molecular Formula      | Number of Atoms of First | Number of Atoms of Second | Name of Compound         |
|------------------------|--------------------------|---------------------------|--------------------------|
| $\text{ClF}$           | 1                        | 1                         | Chlorine monofluoride    |
| $\text{ClF}_5$         | 1                        | 5                         | Chlorine pentafluoride   |
| $\text{CO}$            | 1                        | 1                         | Carbon monoxide          |
| $\text{CO}_2$          | 1                        | 5                         | Carbon dioxide           |
| $\text{Cl}_2\text{O}$  | 2                        | 1                         | Dichlorine monoxide      |
| $\text{PCl}_5$         | 1                        | 5                         | Phosphorus pentachloride |
| $\text{N}_2\text{O}_5$ | 2                        | 5                         | Dinitrogen pentoxide     |

1. Fill in the table to indicate the number of atoms of each type in the molecular formula.
2. Examine the molecular formulas given in Model 1 for various molecular compounds.

a. How many different *elements* are present in each compound shown?

Two

b. Do the compounds combine metals with metals, metals with nonmetals, or nonmetals with nonmetals?

Non-metals with non-metals

c. Based on your answer to b, what type of bonding must be involved in molecular compounds?

Covalent bonds

3. Find all of the compounds in Model 1 that have chlorine and fluorine in them. Explain why the name "chlorine fluoride" is not sufficient to identify a specific compound.

The number of fluorine atoms is different in the two compounds, but name does not reflect the difference

4. Assuming that the name of the compound gives a clue to its molecular formula, predict how many atoms each of these prefixes indicates, and provide two examples.

# Pogil Answers Naming Molecular Compounds

**JR Anderson**



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