

POGIL 2 Ions

How are ions made from neutral atoms?

You have learned that not all atoms of an element are the same. Variation in the number of neutrons results in different isotopes of the element. In this activity we will explore another variation that can take place—the loss and gain of electrons. The exchange of electrons between atoms is a very common way for chemical change to take place. We will see it many times throughout the year.

1. Use Diagram 1 on the next page to complete the following table.

	Metal or Nonmetal	Number of protons in the atom	Number of protons in the ion	Number of neutrons in the atom	Number of neutrons in the ion	Number of electrons in the atom	Number of electrons in the ion	Charge on the ion
Lithium	metal	3	3	4	4	3	2	1+
Magnesium	<i>Metal</i>	12	12	12	12	12	10	2+
Aluminum	<i>Metalloid</i>	13	13	14	14	13	10	3+
Fluorine	<i>Non metal</i>	9	9	10	10	9	10	1-
Oxygen	nonmetal	8	8	8	8	8	10	2-
Nitrogen	<i>Non metal</i>	7	7	7	7	7	10	3-

2. Based on the table you completed in Question 1, what distinguishes a neutral atom from an ion?

A neutral atom has the same number of protons and electrons (no charge). An ion has a charge (# of protons does not equal # of electrons).

3. Examine the isotope symbols in Diagram 1.

- a. Where is the ion charge located in the isotope symbol?

Top right corner

- b. Is a charge indicated on the neutral atoms? If yes, where is it located?

No

4. Which subatomic particle carries a positive charge?

Protons are positive

5. Which subatomic particle carries a negative charge?

Electrons are negative

6. Propose a mathematical equation to calculate the charge on an ion from the number of protons and electrons in an ion. Confirm that your equation works using two positive ion examples and two negative ion examples from Diagram 1.

The charge of an ion = # of protons - # of electrons

Pogil Activities For High School Chemistry Ions Pogil Key

Katrin Zwirgmaier



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Table of Contents Pogil Activities For High School Chemistry Ions Pogil Key

1. Understanding the eBook Pogil Activities For High School Chemistry Ions Pogil Key
 - The Rise of Digital Reading Pogil Activities For High School Chemistry Ions Pogil Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Activities For High School Chemistry Ions Pogil Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pogil Activities For High School Chemistry Ions Pogil Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Activities For High School Chemistry Ions Pogil Key
 - Personalized Recommendations
 - Pogil Activities For High School Chemistry Ions Pogil Key User Reviews and Ratings
 - Pogil Activities For High School Chemistry Ions Pogil Key and Bestseller Lists
5. Accessing Pogil Activities For High School Chemistry Ions Pogil Key Free and Paid eBooks
 - Pogil Activities For High School Chemistry Ions Pogil Key Public Domain eBooks
 - Pogil Activities For High School Chemistry Ions Pogil Key eBook Subscription Services
 - Pogil Activities For High School Chemistry Ions Pogil Key Budget-Friendly Options
6. Navigating Pogil Activities For High School Chemistry Ions Pogil Key eBook Formats

- ePub, PDF, MOBI, and More
 - Pogil Activities For High School Chemistry Ions Pogil Key Compatibility with Devices
 - Pogil Activities For High School Chemistry Ions Pogil Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pogil Activities For High School Chemistry Ions Pogil Key
 - Highlighting and Note-Taking Pogil Activities For High School Chemistry Ions Pogil Key
 - Interactive Elements Pogil Activities For High School Chemistry Ions Pogil Key
 8. Staying Engaged with Pogil Activities For High School Chemistry Ions Pogil Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pogil Activities For High School Chemistry Ions Pogil Key
 9. Balancing eBooks and Physical Books Pogil Activities For High School Chemistry Ions Pogil Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pogil Activities For High School Chemistry Ions Pogil Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Pogil Activities For High School Chemistry Ions Pogil Key
 - Setting Reading Goals Pogil Activities For High School Chemistry Ions Pogil Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Pogil Activities For High School Chemistry Ions Pogil Key
 - Fact-Checking eBook Content of Pogil Activities For High School Chemistry Ions Pogil Key
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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