

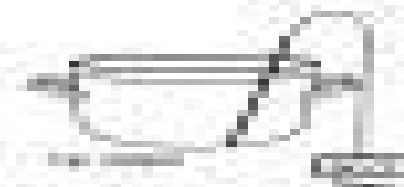
Calorimetry

What is calorimetry? How does it work? What is its purpose?

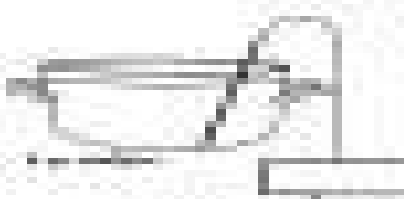
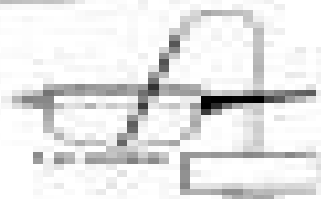
When a substance is heated, the temperature of that substance increases. Will the amount of energy (heat) required to heat different substances to the same temperature increase? Will the same amount of energy be needed to heat different temperatures? Increase or decrease amounts of the same substance to different temperatures will require less mass, temperature, heat energy, and the type of substance are related.

Model 1 – A Plot of Mass

Calorimetry



The energy is constant at 1000 J for all cases.



1. In Model 1, which substance has the more mass of heat? Substance A

2. Compare the graphs above as Model 1.

a. How many (what units) of energy were added to the substance? 1000 J

b. How many (what units) of energy were added to the substance? 1000 J

c. In Model 1, which substance has the more mass of heat? (What units) Substance A

Why? (What is the reason for this?) Substance A has a higher temperature than Substance B.

Pogil Activities For High School Chemistry Calorimetry Packet

David Kirk



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