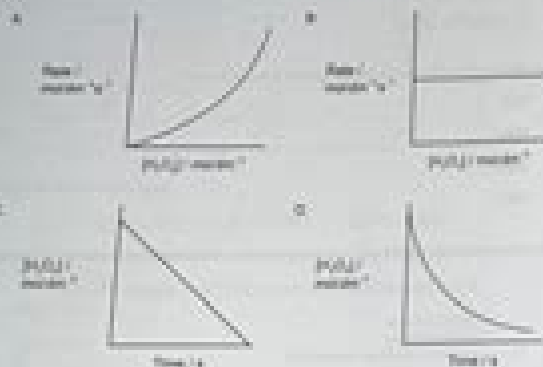


12. Decomposition of hydrogen peroxide in an aqueous solution proceeds as follows:



The rate expression for the reaction was found to be: rate = $k[\text{H}_2\text{O}_2]$

Which graph is consistent with the given rate expression?



13. The rate constant, k , is commonly described by the Arrhenius equation: $k = Ae^{-E_a/RT}$. Which of the following statements are correct?

- A greater E_a value results in a smaller k value.
- Reactions of less complex molecules usually have a greater value of A .
- The slope (gradient) of $\ln k$ versus $1/T$ equals E_a .

- A. i and ii only
B. i and iii only
C. ii and iii only
D. i, ii and iii

Turn over

14. What happens when the temperature of the following equilibrium system is increased?



Position of equilibrium	Equilibrium value of forward and reverse rate constants
A. shifts to the left	increase
B. shifts to the left	decrease
C. shifts to the right	decrease
D. shifts to the right	increase

15. A mixture of 0.40 mol of CO(g) and 0.40 mol of $\text{H}_2(\text{g})$ was sealed in a 1.00 dm³ vessel. The following equilibrium was established:



At equilibrium, the mixture contained 0.20 mol of CO(g) . How many moles of $\text{H}_2(\text{g})$ and $\text{CH}_3\text{OH(g)}$ were present at equilibrium?

	Equilibrium mol of H_2	Equilibrium mol of CH_3OH
A.	0.20	0.20
B.	0.50	0.20
C.	0.20	0.20
D.	0.20	0.75

16. Which species behave as Brønsted-Lowry bases in the following reaction?



- A. HNO_3 and HNO_3
B. HNO_3 and H_2SO_4
C. H_2SO_4 and HNO_3
D. H_2SO_4 and HNO_3

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