| - | Using the redox systems in Table 6.1, construct an equation for a reaction between acidific | e |
|---|---|---|
| | dichromate(VI) ions and methanoic acid, HCOOH. | |

Rather than using [O] or [H], your equation must show the actual reactants and products.

| | | [2] |
|-----|-------|--|
| (d) | | tudent added some chromium metal to an acidified solution containing copper(II) ions. A ction took place. The student concluded that 'chromium is more reactive than copper'. |
| | (i) | Explain, in terms of their electrode potentials, why 'chromium is more reactive than copper' in this reaction. |
| | | |
| | | |
| | | *************************************** |
| | | [2] |
| | (ii) | When this experiment was carried out, the student observed some bubbles of a gas. |
| | | Suggest an explanation for this observation. |
| | | |
| | | [1] |
| (e) | Cel | thanoic acid, HCOOH, has the common name of 'formic acid'. Direct-Formic Acid Fuel Is (DFAFCs) are being developed for use in small, portable electronics such as phones I laptop computers. |
| | In ti | his fuel cell, methanoic acid (the fuel) reacts with oxygen to generate a cell potential. |
| | (i) | Predict the standard cell potential of a DFAFC. |
| | | standard cell potential =V [1] |
| | (ii) | Suggest two advantages of using methanoic acid as the fuel in a fuel cell rather than hydrogen. |
| | | |

[Total: 14]

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