

Question number	Answer	Marks	Guidance
1	$\text{Ca} : \text{N} : \text{O} = 30.35/40.1 : 21.20/14.0 : 48.45/16.0$ $= 0.7569 : 1.5286 : 3.028$ Formula = CaN_2O_4	B1 B1	
2 (a)	$2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$	B1	
2 (b)	$4.50 \times 10^{-3} \text{ mol}$	B1	
2 (c)	$n(\text{KClO}_3) = 3.0 \times 10^{-3} \text{ mol}$ $M(\text{KClO}_3) = 122.6 \text{ g mol}^{-1}$ mass of $\text{KClO}_3 = 0.3678 \text{ g}$	B1 B1 B1	
3	$M(\text{ZnSO}_4) = 161.5 \text{ g mol}^{-1}$ $n(\text{ZnSO}_4) = 6.57 \times 10^{-3} \text{ mol}$ $n(\text{H}_2\text{O}) = (1.893 - 1.061)/18 = 4.62 \times 10^{-2} \text{ mol}$ $x = n(\text{H}_2\text{O})/n(\text{ZnSO}_4) = 7$ Formula = $\text{ZnSO}_4 \cdot 7\text{H}_2\text{O}$	B1 B1 B1 B1	
4	$M(\text{Na}_2\text{CO}_3) = 106.0 \text{ g mol}^{-1}$ Actual $n(\text{Na}_2\text{CO}_3) = 0.0234 \text{ mol}$ Theoretical $n(\text{Na}_2\text{CO}_3) = 0.0234 \times 100/65$ $= 0.0360 \text{ mol}$ Theoretical $n(\text{NaHCO}_3) = 0.0720 \text{ mol}$ Mass of $\text{NaHCO}_3 = 0.0720 \times 84.0 = 6.05 \text{ g}$	B1 B1 B1 B1 B1	
5 (a)	$\text{Fe}_2\text{O}_3(\text{s}) + 3\text{CO}(\text{g}) \rightarrow 2\text{Fe}(\text{s}) + 3\text{CO}_2(\text{g})$	B1	
5 (b)	$M(\text{Fe}_2\text{O}_3) = 159.6 \text{ g mol}^{-1}$ $n(\text{Fe}_2\text{O}_3) = (10\,000 \times 10^6)/159.6 = 6.266 \times 10^7 \text{ mol}$ $n(\text{Fe}) = 1.253 \times 10^8 \text{ mol} = 6.992 \times 10^5 \text{ g}$ (6992 tonnes)	B1 B1 B1	
6 (a)	$\text{Na} : \text{N} : \text{O} = 27.1/23.0 : 16.5/14.0 : 56.4/16.0$ $= 1.178 : 1.179 : 3.525$ Formula = NaNO_3	B1 B1	
6 (b)	$2\text{NaNO}_3(\text{s}) \rightarrow 2\text{NaNO}_2(\text{s}) + \text{O}_2(\text{g})$	B1	

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SJ Ball



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