

16. The dissociation of water is shown below:



- At 25 °C, the auto product of water,  $K_w$ , is  $1.0 \times 10^{-14}$  mol/L.  
At 65 °C, the auto product of water,  $K_w$ , is  $1.0 \times 10^{-13}$  mol/L.

17. Explain whether the dissociation of water is an endothermic or exothermic process.  
Endothermic

18. Predict, using a calculation, whether a solution of 1.0 M  $HCl$  is neutral, acidic or basic. (2)

Acidic

19. Explain by calculating pH.

20.  $pK_a$ ,  $pH$  and  $pOH$  are significant notions.

Calculate  $pK_a$  of  $HCl$ .

Give your answer to two decimal places.

$$-\log(4.31 \times 10^{-4}) = 3.07$$

$$pK_a = 3.07$$

21. A solution of hydrochloric acid is mixed with water and the resulting solution is neutral at 25 °C.

Calculate the pH of the resulting solution at 65 °C.

Give your answer to two decimal places.

$$pH = 13.01$$

pH = \_\_\_\_\_ (2)

(Total: 10)

# Ocr Chemistry A2 Unofficial Mark Scheme 2014

**L Cohen**



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