

## Chemistry – Gas Laws ACT Questions [USE TO STUDY FOR YOUR TEST]

### Multiple Choice

Identify the choice that best completes the statement or answers the question.

1. At  $22^{\circ}\text{C}$ , the air pressure in a car tire is 1293 torr. Convert this pressure to atmospheres (atm).
- 12.76
  - 5.262
  - 1.701
  - 1.293

2. This diagram shows a closed-end mercury (Hg) barometer.



When atmospheric pressure is 1.30 atm, what is the height ( $h$ ), in millimeters, of the Hg in the barometer?

- 132
  - 585
  - 760
  - 988
3. A chemist compresses a 1.5 L sample of carbon dioxide ( $\text{CO}_2$ ) gas at 1.2 atm to a final volume of 0.75 L. Assuming the temperature is constant, what is the final pressure, in atm, exerted by the  $\text{CO}_2$  gas?
- 0.60
  - 0.94
  - 1.7
  - 2.4
4. A vessel connected to an open-end mercury (Hg) manometer contains nitrogen ( $\text{N}_2$ ) gas. The atmospheric pressure is 0.993 atm. The height difference between the two arms of the manometer



is 122 mm.

What is the pressure, in mm Hg, of the  $\text{N}_2$  gas?

- 355
- 633

- 755
- 877

5. At 1.00 atm, a sealed weather balloon contains 20.0 L of helium (He) gas at  $25.0^{\circ}\text{C}$ . Assume that none of the He escapes and the pressure is constant. What is the volume, in liters, of He in the weather balloon at  $35.0^{\circ}\text{C}$ ?

- 40.0
- 28.0
- 20.7
- 14.3

6. At 1254 m above sea level, the city of Helena, Montana, has an average atmospheric pressure of 0.859 atm. A chemistry student in Helena finds that pure distilled water boils at  $95.9^{\circ}\text{C}$ . Which statement explains the student's results at this high elevation?

- The boiling point of water increases due to lower atmospheric pressure.
- The boiling point of water increases due to higher atmospheric pressure.
- The boiling point of water decreases due to lower atmospheric pressure.
- The boiling point of water decreases due to higher atmospheric pressure.

7. For a chemistry homework assignment, Beth must consider the effusion rates of methane ( $\text{CH}_4$ ) gas and nitrogen trifluoride ( $\text{NF}_3$ ) gas at  $25^{\circ}\text{C}$ . Effusion occurs as gas particles escape from a container through a small hole. Which statement accurately compares the effusion rates of  $\text{CH}_4$  and  $\text{NF}_3$  at  $25^{\circ}\text{C}$ ?

- $\text{CH}_4$  effuses 2 times faster than  $\text{NF}_3$ .
- $\text{CH}_4$  effuses 4 times faster than  $\text{NF}_3$ .
- $\text{NF}_3$  effuses 2 times faster than  $\text{CH}_4$ .
- $\text{NF}_3$  effuses 4 times faster than  $\text{CH}_4$ .

8. At standard temperature and pressure, a chemist burns 3.50 L of pentane ( $\text{C}_5\text{H}_{12}$ ) in excess oxygen ( $\text{O}_2$ ) gas.
- $$\text{C}_5\text{H}_{12}(l) + 8 \text{O}_2(g) \rightarrow 5 \text{CO}_2(g) + 6 \text{H}_2\text{O}(l)$$
- What is the maximum number of liters of carbon dioxide ( $\text{CO}_2$ ) that the reaction can produce?

- 17.5
- 8.50
- 5.60
- 0.700

# Sample Test Questions On Gas Laws

**J Dewey**



## Sample Test Questions On Gas Laws:

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