

# Physical Science – 2nd Semester – Final Exam Study Guide – (2012–2013)

## 1. Physical Science Basics

- What tool(s) would you want to use to find the:
  - Mass of an object? triple beam balance Basic SI Unit? kg
  - Volume of an object? ruler or water displacement method Basic SI Unit? cm<sup>3</sup> or ml
  - Length of an object? Ruler Basic SI Unit? meter
  - Density of an object? graduated cylinder Basic SI Unit? ml
- What is the formula for calculating density? mass/volume
  - What is the density of a cube that is 25 g and has a volume of 3 cm<sup>3</sup>? 8g/cm<sup>3</sup>
  - Compare the density of solid H<sub>2</sub>O to liquid H<sub>2</sub>O, and explain what observable effect this has.  
solid H<sub>2</sub>O less dense than liquid H<sub>2</sub>O

## 2. Chemistry Basics

- Element (def.) substance made up of one kind of atom
- Atom (def.) basic building block of matter
- Compound (def.) a molecule that contains atoms or molecules of two or more different elements
- Molecule (def.) a combination of two or more of the same or different atoms that are bonded together

## 3. Model of the Atomic Evolution of the Atomic Theory

Atoms are so small that many scientists throughout history have created models to describe them.

Why have atomic theories changed over time? Better technology came around.

| Scientist         | Year | Major Discovery/ Contribution                                                                                                                                                                                           | Name of atomic model |
|-------------------|------|-------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|----------------------|
| John Dalton       | 1808 | Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different masses.                                                                                                | Billiard Ball Model  |
| J.J. Thomson      | 1897 | Negatively charged electron                                                                                                                                                                                             | Plum Pudding Model   |
| Niels Bohr        | 1913 | Electrons move <u>fixed orbits</u> around the atom's nucleus in specific layers, or <u>electron shells</u> . Atoms absorb or give off energy when the electrons move from one shell ( <u>energy level</u> ) to another. | Bohr Model           |
| Ernest Rutherford | 1909 | 1909 /Based on idea that positive reflects positive/ atom mostly empty space; electrons orbit randomly around positively charged nucleus                                                                                | Planetary Model      |

## 4. Periodic Table

- The current periodic table is arranged by increasing atomic number.
- Reading the Periodic Table
  - An element's properties can be predicted from its location in the periodic table.
  - The atomic number represents the number of protons.
  - The atomic mass represents the number of protons and neutrons.

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