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4.14.

$$\left(\frac{\partial H_m}{\partial T}\right)_P = \frac{C_{P,m}}{T} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1}$$

$$\text{and } C_{P,m} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1} = 1.25 \left(\frac{1000 \text{ J}}{1 \text{ kJ}} \right) = 1.25 \text{ kJ K}^{-1} \text{ mol}^{-1}$$

and that we can estimate $C_{V,m}$ as follows:

- 4.15. (a) The noble gas with the fewest atoms has the smallest percentage of light.
 (b) The noble gas with the greatest number of atoms has the greatest percentage of light.
 (c) The noble gas with the greatest number of atoms has the smallest percentage of light.

4.16. From the data we find the heat capacity of the gas is $C_{V,m} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1}$, and that the heat capacity is $C_{P,m} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1}$.

$$\text{or } C_{V,m} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1} = 1.25 \left(\frac{1000 \text{ J}}{1 \text{ kJ}} \right) = 1.25 \text{ kJ K}^{-1} \text{ mol}^{-1}$$

4.17. From the data we find $C_{V,m} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1}$.

$$\text{and } C_{P,m} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1} = 1.25 \left(\frac{1000 \text{ J}}{1 \text{ kJ}} \right) = 1.25 \text{ kJ K}^{-1} \text{ mol}^{-1}$$

$$\text{and } C_{V,m} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1} = 1.25 \left(\frac{1000 \text{ J}}{1 \text{ kJ}} \right) = 1.25 \text{ kJ K}^{-1} \text{ mol}^{-1}$$

4.18. From the data we find the heat capacity of the gas is $C_{V,m} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1}$.

$$\text{or } C_{V,m} = 1.25 \text{ J K}^{-1} \text{ mol}^{-1} = 1.25 \left(\frac{1000 \text{ J}}{1 \text{ kJ}} \right) = 1.25 \text{ kJ K}^{-1} \text{ mol}^{-1}$$

Physical Chemistry Ball Solution Manual

**Hans Kuhn, David H. Waldeck, Horst-
Dieter Försterling**



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