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# Chapter 7

# Equilibria for Multiple Component Systems

- 7.2. With only a single component and phase, if only Fe<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> were present in our system, we would have C-P+2=1-1+2=2 degrees of freedom.
- 7.4. For a single-component system, there will never be negative degrees of freedom. The maximum number of phases that can exist in equilibrium is three, so the Gibbs phase rule would yield 1 3 + 2 = 0 degrees of freedom (like might occur at the triple point). For a two-component system, the maximum number of phases that can exist in equilibrium is four. There will never be negative degrees of freedom.
- 7.6. In this reaction, there are only three independent components (the amount of the fourth can be determined from the equilibrium constant) and 3 phases, two solids and a gas mixture. Therefore, there are 3 3 + 2 = 2 degrees of freedom.
- 7.8. For water:  $n = \frac{pV}{RT} = \frac{(23.76 \text{ torr})(5.00 \text{ L})}{(0.08205 \text{ L} \cdot \text{atm/mol} \cdot \text{K})(298.15 \text{ K})} \times \frac{1 \text{ atm}}{760 \text{ torr}} = 6.39 \times 10^{-3} \text{ mol of water is needed to ensure that there is a liquid and gas phase. This equals 0.115 grams. For methanol:$

$$n = \frac{pV}{RT} = \frac{(125.0 \text{ torr})(5.00 \text{ L})}{(0.08205 \text{ L} \cdot \text{atm/mol} \cdot \text{K})(298.15 \text{ K})} \times \frac{1 \text{ atm}}{760 \text{ torr}} = 3.36 \times 10^{-2} \text{ mol of methanol is}$$
needed to ensure that there is a liquid and gas phase. This equals 1.08 grams.

7.10. Since, according to equation 7.11,  $a_i = \frac{p_i}{p_i}$ , and since  $p_i^* = 760$  torr, we can calculate  $a_i$ :

# **Physical Chemistry Ball Instructor Manual**

**CO Houle** 

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#### **Table of Contents Physical Chemistry Ball Instructor Manual**

- 1. Understanding the eBook Physical Chemistry Ball Instructor Manual
  - The Rise of Digital Reading Physical Chemistry Ball Instructor Manual
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Physical Chemistry Ball Instructor Manual
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Physical Chemistry Ball Instructor Manual
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Physical Chemistry Ball Instructor Manual
  - Personalized Recommendations

- Physical Chemistry Ball Instructor Manual User Reviews and Ratings
- Physical Chemistry Ball Instructor Manual and Bestseller Lists
- 5. Accessing Physical Chemistry Ball Instructor Manual Free and Paid eBooks
  - Physical Chemistry Ball Instructor Manual Public Domain eBooks
  - Physical Chemistry Ball Instructor Manual eBook Subscription Services
  - Physical Chemistry Ball Instructor Manual Budget-Friendly Options
- 6. Navigating Physical Chemistry Ball Instructor Manual eBook Formats
  - o ePub, PDF, MOBI, and More
  - Physical Chemistry Ball Instructor Manual Compatibility with Devices
  - Physical Chemistry Ball Instructor Manual Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - o Adjustable Fonts and Text Sizes of Physical Chemistry Ball Instructor Manual
  - o Highlighting and Note-Taking Physical Chemistry Ball Instructor Manual
  - Interactive Elements Physical Chemistry Ball Instructor Manual
- 8. Staying Engaged with Physical Chemistry Ball Instructor Manual
  - o Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Physical Chemistry Ball Instructor Manual
- 9. Balancing eBooks and Physical Books Physical Chemistry Ball Instructor Manual
  - ∘ Benefits of a Digital Library
  - Creating a Diverse Reading Collection Physical Chemistry Ball Instructor Manual
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Chemistry Ball Instructor Manual
  - Setting Reading Goals Physical Chemistry Ball Instructor Manual
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Chemistry Ball Instructor Manual
  - Fact-Checking eBook Content of Physical Chemistry Ball Instructor Manual

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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