

Physical Science – 2nd Semester – Final Exam Study Guide – (2012–2013)

1. Physical Science Basics

- What tool(s) would you want to use to find the:
 - Mass of an object? triple beam balance
 - Volume of an object? ruler or water displacement method
 - Length of an object? Ruler
 - Density of an object? graduated cylinder
- What is the formula for calculating density? mass/volume
 - What is the density of a cube that is 25 g and has a volume of 3 cm³? 8.33/cm³
 - Compare the density of solid H₂O to liquid H₂O, and explain what observable effect this has.
solid H₂O less dense than liquid H₂O

2. Chemistry Basics

- Element (def.) substance made up of one kind of atom
- Atom (def.) basic building block of matter
- Compound (def.) a molecule that contains atoms or molecules of two or more different elements
- Molecule (def.) a combination of two or more of the same or different atoms that are bonded together

3. Model of the Atomic Evolution of the Atomic Theory

Atoms are so small that many scientists throughout history have created models to describe them.

Why have atomic theories changed over time? Better technology came around.

Scientist	Year	Major Discovery/ Contribution	Name of atomic model
John Dalton	1808	Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different masses.	Billiard Ball Model
J.J. Thomson	1897	Negatively charged electron	Plum Pudding Model
Niels Bohr	1913	Electrons move <u>fixed orbits</u> around the atom's nucleus in specific layers, or <u>electron shells</u> . Atoms absorb or give off energy when the electrons move from one shell (<u>energy level</u>) to another.	Bohr Model
Ernest Rutherford	1909	1909 /Based on idea that positive reflects positive/ atom mostly empty space; electrons orbit randomly around positively charged nucleus	Planetary Model

4. Periodic Table

- The current periodic table is arranged by increasing atomic number.
- Reading the Periodic Table
 - An element's properties can be predicted from its location in the periodic table.
 - The atomic number represents the number of protons.
 - The atomic mass represents the number of protons and neutrons.

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systems are crucial to public safety for years the designers of those systems had few published resources to reference and guide them through their design processes The first edition of this book changed all that and now The Design and Layout of Fire Sprinkler Systems Second Edition suits their needs even better Written and thoroughly updated by a fire prevention engineer with more than 20 years of experience this book provides a complete systematic introduction to automatic fire sprinkler design and layout from design basics code requirements and pipe hanging to hydraulic calculations retrofits and details on fire pumps The author carefully outlines all of a designer s responsibilities and includes an entire chapter dedicated to preparing for the NICET exam More than 150 sample diagrams checklists sample forms spec sheets photographs and a glossary complement the text and the larger page size of this edition permits clear presentation of diagrams and schematics The Design and Layout of Fire Sprinkler Systems not only builds the foundation and skills of newcomers to the field but also provides an outstanding reference for fire safety professionals building inspectors insurance underwriters and municipal officials **Chemical News and Journal of Physical Science** ,1895 *Physics Handbook & Study Guide Grade 11-12 IEB* Kevin Smith,2024-02-01 A comprehensive summary of Grade 11 12 Physics Simple logical

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