



Name \_\_\_\_\_ Date \_\_\_\_\_ Period \_\_\_\_\_

### GRAPHING LOGARITHMIC FUNCTIONS WORKSHEET

Transformations of Logarithmic Functions:  $y = a \log_b(x - h) + k$ , where  $a$  is the vertical stretch or shrink,  $h$  is the horizontal translation and  $k$  is the vertical translation. The parent graph  $y = \log_b x$  passes through the points  $(1, 0)$  and  $(b, 1)$  and has a vertical asymptote at  $x = 0$ .

Match the function with its graph.

1.  $f(x) = \log_2 x$

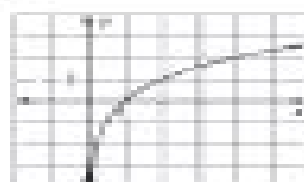
2.  $f(x) = \log_2 x$

3.  $f(x) = \log_{1/2} x$

A.



B.



C.

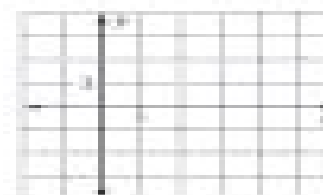
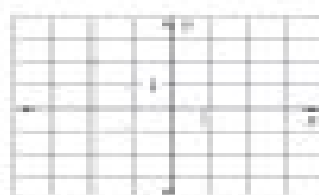
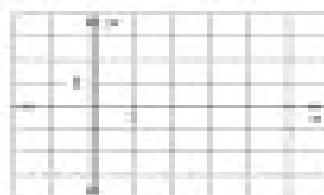


Graph the function. State the domain and range. Identify the parent function and describe the transformations.

4.  $f(x) = \log_2 x$

5.  $f(x) = \log_2(x + 2)$

6.  $f(x) = -\log_2 x - 1$



# Rpdp Graphing Logarithmic Functions Worksheet Answers

**Karin Nielsen-Saines**



## **Rpdp Graphing Logarithmic Functions Worksheet Answers:**

## The Enigmatic Realm of **Rpdp Graphing Logarithmic Functions Worksheet Answers**: Unleashing the Language is Inner Magic

In a fast-paced digital era where connections and knowledge intertwine, the enigmatic realm of language reveals its inherent magic. Its capacity to stir emotions, ignite contemplation, and catalyze profound transformations is nothing short of extraordinary. Within the captivating pages of **Rpdp Graphing Logarithmic Functions Worksheet Answers** a literary masterpiece penned by a renowned author, readers attempt a transformative journey, unlocking the secrets and untapped potential embedded within each word. In this evaluation, we shall explore the book's core themes, assess its distinct writing style, and delve into its lasting effect on the hearts and minds of people who partake in its reading experience.

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web solution 1 the two half reactions balanced as if in acidic solution  $2\text{CN}^{-} + \text{Au} + \text{CN}^{-} + 2\text{e}^{-} + 2\text{H}^{+} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{HCN}$  make electrons equal convert to basic solution  $4\text{CN}^{-} + 2\text{Au} + 2\text{Au} + \text{CN}^{-} + 2\text{e}^{-}$  multiplied by a factor of 2  $2\text{e}^{-} + 2\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{OH}^{-}$  the final answer

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web bonus problem here is a balanced net ionic equation 3br<sup>-</sup> 2 6oh<sup>-</sup> 5br<sup>-</sup> bro<sup>-</sup> 3 3h<sup>+</sup> 2 o the above equation takes place in two stages first bromine reacts with oh ions to form bromide ions and bro ions second the bro ions react to form bromide ions and bromate v ions

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation ni s ni a 2 a q 2 e now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balance each half reaction in basic solution 4 5 6 7  $\text{Cr}_2\text{O}_7^{2-}$   $\text{NO}_3^-$   $\text{SO}_4^{2-}$   $\text{MnO}_2$   $\text{Cr}_3$   $\text{NO}_3$   $\text{SO}_2$   $\text{Mn}_2\text{O}_3$  balance each redox reaction in acid solution using the half reaction method 8  $\text{H}_2\text{O}_2$   $\text{Cr}_2\text{O}_7^{2-}$  2 9  $\text{TeO}_3$  2  $\text{N}_2\text{O}_4$   $\text{O}_2$   $\text{Cr}_3$   $\text{Te}$  10  $\text{ReO}_4$  11  $\text{PbO}_2$   $\text{IO}_2$   $\text{Pb}_2$  12  $\text{As}_2\text{O}_3$   $\text{IO}_3^-$   $\text{Re}$   $\text{IO}_3^-$   $\text{H}_2\text{AsO}_4$

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively  $\text{C}_6\text{H}_8\text{O}_6$   $2\text{H}_2\text{NO}_2$   $\text{C}_6\text{H}_6\text{O}_6$   $2\text{H}_2\text{O}_2$   $\text{NO}$  choose 1 answer  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{NO}_2$  a  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{NO}_2$  b  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{NO}_2$  c  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{NO}_2$  d  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{NO}_2$

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