



NATIONAL  
SENIOR CERTIFICATE

GRADE 11

NOVEMBER 2020

PHYSICAL SCIENCES P2  
(CHEMISTRY)  
EXEMPLAR

MARKS: 100  
TIME: 1 hour



The question paper consists of 10 pages, including 4 extra sheets.  
Use blue or black ink for writing. Do not use correction fluid.

# QUESTION 3

## PHYSICAL SCIENCES PAPER 2

## INTERMOLECULAR FORCES

**GRADE 11**

**NOVEMBER 2020**

Examiner's use: \_\_\_\_\_ Physical Sciences P2

QUESTION 3 starts on a NEW page.

3.1 The boiling points in the table below were obtained during an investigation carried out to measure the boiling points of pure 1-hexanol. The table also lists the  $\Delta H_{\text{vap}}$  and  $\Delta H_{\text{fusion}}$  for 1-hexanol in kJ mol<sup>-1</sup> and 1 mol of 1-hexanol.

| Compound | Molecular mass (M <sub>r</sub> ) | Boiling point (°C) |
|----------|----------------------------------|--------------------|
| A        | 102                              | 16,2               |
| B        | 102                              | 21                 |

- 3.1.1 Define the term boiling point. (2)
- 3.1.2 Is pure 1-hexanol only composed of A and B or of A, B and C or of others as well? (2)
- 3.1.3 Name the type of intermolecular force that exists between molecules of both compounds A and B and is the prime reason of their volatility. (2)
- 3.1.4 What factor of the compounds A and B is responsible for the difference in their boiling points? (2)
- 3.1.5 What compound (A or B) will have the higher vapour pressure? (2)
- 3.1.6 Which compound (A or B) will have the higher surface tension? (2)

Compound C (M<sub>r</sub> 102) will also be in the table. Use a suitable intermolecular force that both compounds A and B do not have as a secondary factor during point of 3.1.2.

- 3.1.8 Explain why the boiling point of compound B is 100°C higher than that of compound A and B for which the boiling point is 21°C, and intermolecular forces are involved. (2)

3.2 Add column C to the table in Table 1 below. It contains compounds C, D, E and F. It is a mixture of pure substances in the liquid state.

|     |     |     |     |     |     |
|-----|-----|-----|-----|-----|-----|
| 102 | 102 | 102 | 102 | 102 | 102 |
|-----|-----|-----|-----|-----|-----|

- 3.2.1 Is there any difference in the boiling points of the substances? (2)
- 3.2.2 Explain the reason for QUESTION 3.2.1 about the volatility in the intermolecular forces involved. (2)

Examiner's use

Page 10 of 10

# Physical Science Paper 2 Intermolecular Forces Investigation

**Daniel F McAuley**



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## **Table of Contents Physical Science Paper 2 Intermolecular Forces Investigation**

1. Understanding the eBook Physical Science Paper 2 Intermolecular Forces Investigation
  - The Rise of Digital Reading Physical Science Paper 2 Intermolecular Forces Investigation
  - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Paper 2 Intermolecular Forces Investigation
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Physical Science Paper 2 Intermolecular Forces Investigation
  - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Paper 2 Intermolecular Forces Investigation
  - Personalized Recommendations
  - Physical Science Paper 2 Intermolecular Forces Investigation User Reviews and Ratings
  - Physical Science Paper 2 Intermolecular Forces Investigation and Bestseller Lists

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  - Physical Science Paper 2 Intermolecular Forces Investigation Enhanced eBook Features
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  - Highlighting and Note-Taking Physical Science Paper 2 Intermolecular Forces Investigation
  - Interactive Elements Physical Science Paper 2 Intermolecular Forces Investigation
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  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Physical Science Paper 2 Intermolecular Forces Investigation
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  - Creating a Diverse Reading Collection Physical Science Paper 2 Intermolecular Forces Investigation
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  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Physical Science Paper 2 Intermolecular Forces Investigation
  - Setting Reading Goals Physical Science Paper 2 Intermolecular Forces Investigation
  - Carving Out Dedicated Reading Time
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  - Fact-Checking eBook Content of Physical Science Paper 2 Intermolecular Forces Investigation
  - Distinguishing Credible Sources
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- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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