A Company of the Comp	Professional	District

#### Physical Science — 2nd Semester — Final Exam Study Guide — (2012-2013)

#### Physical Science Basics

What tooks would you want to use to find the:

Mass of an object? triple beam balance Barrier Strictory No.

Volume of an object? rules or water displacement method Restic St Unit? and or red.

Length of an object? Ruler

Stack: S1 Unit? meter Density of an object? graduated cylinder Blooks St Unit 7 mil.

What is the formula for calculating density? resun/volume.

What is the density of a cube that is 35 g and has a volume of 3 cm 7 Se/cm

Compare the density of solid NLO to liquid NLO, and explain what observable effect this has.

solid HyDis iess dense than liquid HyD.

#### Chemistry Basics.

Dernet (def.) substance made up of one kind of store

Atom (def.) basic building block of matter.

Composed (def.) a majerate that centains atoms or molecules of two or more different elements

4. Molecule (4ef.) a combination of two at more of the same or different atoms that are bonded together

#### Model of the Atomic Evolution of the Atomic Theory

Atoms are so small that reary scientists throughout history have created anodels to describe them.

Why have storaic theories channel over time? Better technology come around.

Solumber	Year	Major Discovery/ Contribution	Morne of atomic model
John Deltse	1862	Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different mesors.	Billiard Boll Model
J.J. Thoreson	1887	Negatively charged electron	Plum Pudding Model
Niets Bohr	1909	Electrons move fined arbits around the atom's nucleus in specific layers, or siectron shalls. Atoms absorb or give off energy when the electrons move from one shall (grappy level) to enother	Stohr Model
orwest Muldherford	1900	1909 (Based on lifes that positive reflects positive) atom mostly empty space; electrons orbit rendemly around positively charged reciess	Planetary Model

#### 4. Previouslie Testates

- The current periodic table is arranged by increasing atomic number.
- Reackes the Periodic Table
  - An element's properties can be predicted from its location in the periodic table.
  - The atomic number represents the number of pretent.
  - The atomic mass represents the number of pentrons and posture

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**Christian Drosten** 

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