

Physical Science Standards:

SPS1: Students will investigate our current understanding of the atom.

a. Examine the structure of the atom in terms of

- proton, electron, and neutron locations.
 - Proton has a charge of _____ and is located in the _____ of the atom.
 - Protons are also the elements _____ number of the identity of the element, so protons are found by looking at the atomic number
 - Neutron has a charge of _____ and is located in the _____ of the atom.
 - Neutrons are found by subtracting the _____ Number from the _____ Mass or the BIG #- SMALL #.
 - Electron has a charge of _____ and is located in the _____ of the atom.
 - Electrons are equal to the number of _____ in a neutral atom.
 - The nucleus is made up of these 2 particles, _____ and _____.
 - The electrons are located in this region of the atom, known as the _____.
 - The electron cloud is divided in to energy levels or energy shells, the first level will hold _____ electrons and the second level will hold _____ electrons.
 - The valence shell or valence level is the _____ energy shell of the atom.
- atomic mass and atomic number.
 - Atomic number is equal to the number of _____ in the atom and is also the atoms identity—it's the small number of the periodic table.
 - Atomic mass is equal to the average mass of all the isotopes of a particular element—it the big number on the periodic table.
- atoms with different numbers of neutrons (isotopes).
 - Define Isotope _____.
 - Ex. Carbon-12 and Carbon-14
 - Carbon 12 has _____ P _____ N _____ E
 - Carbon 14 has _____ P _____ N _____ E
- explain the relationship of the proton number to the element's identity.
 - The _____ number is equal to the number of _____ in the atom which is how the element is identified.
 - Ex. Nitrogen has an atom number of _____, which means that is has _____ protons.

molybdenum	← element name
42	← atomic number number of protons (Z)
Mo	← atomic symbol
95.94	← atomic mass A (this is an average mass)



- Number of Protons = Atomic Number
- Number of Electrons = Number of Protons = Atomic Number
- Number of Neutrons = Mass Number - Atomic Number

Physical Science Eoct Study Guide Forsyth

Wolfgang Guggemos



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didn't struggle with the subject they struggled with the language It was never about brains or ability His students simply didn't have the knowledge of the specific language needed to succeed Through experimentation and research he discovered that for any subject there was a list of essential words that when mastered unlocked a student's ability to progress in the subject Lewis called this set of vocabulary the Insider's Words When he applied these Insider's Words the results were incredible His students began to learn with ease He was on his way to developing the landmark series of workbooks and applications to teach this Insider's Language to students around the world

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Success: Master the Key Vocabulary of the Georgia End of Course Physical Science Exam Lewis Morris, 2019-01-07

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