



Pays De La Loire Mayenne

Sabine Zange



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web during this reaction the h^+ ions from the acid and the oh^- ions from the alkali react together to form water as this happens energy is released which is detected by the thermometer

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kj less exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of sodium hydroxide solution with hcl as 57.9 kj mol⁻¹

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eg 1 $\text{hcl aq} + \text{naoh aq} \rightarrow \text{nacl aq} + \text{h}_2\text{o l}$ for which $\Delta h = 57\text{ kJ mol}^{-1}$ eg 2 $\text{h}_2\text{so}_4 \text{ aq} + \text{koh aq} \rightarrow \text{kn o}_3 \text{ aq} + \text{h}_2\text{o l}$ for which $\Delta h = 57\text{ kJ mol}^{-1}$

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