

PRE-AP CHEMISTRY: Semester 2 Final Review

Final: June 8 (period 7); June 9 (period 6)

Review days after school: Monday June 1, Wednesday June 3, Thursday June 4

*You have come a long way this year! You should celebrate your successes and continue to work hard until the end. Let me know how I can help you!

**You can also use a periodic table and your reference sheets from the semester.

***Your binder contains your work from the whole semester (maybe even the whole year!) You can study from your semester's old worksheets, practice problems within the chapters, labs, quizzes, or tests, or look in my folders for extra worksheets/practice.

I. SEMESTER 1:

- a. Significant figures: rules for measuring, rounding, etc.
 - i. Why do we need them?
 - ii. How can you tell how many sig figs are in a measurement?
- b. Scientific notation, converting metric units
 - i. $5.64 \times 10^3 =$ _____
 - ii. $0.000374 =$ _____
 - iii. $5000 \text{ grams} =$ _____ kg
 - iv. $0.345 \text{ L} =$ _____ mL
- c. Physical/chemical changes (in Ch. 9 again)
- d. Structure of an atom: numbers of protons, neutrons, electrons
 - i. What are isotopes?
 - ii. How do you identify the number of protons, neutrons, and electrons in an atom?
- e. Electron configurations
 - i. Use the Aufbau principle to write e^- configs for any element on P-table
 - ii. Know electron configurations for **ions**; what has changed? Why do atoms become ions?
- f. Arrangement of periodic table
 - i. Names of specific groups on table: halogens, alkali metals, noble gases, metals, nonmetals

II. UNIT 5 IONIC BONDING (Chapter 4/12)

- i. How do positive ions form? Negative ions?
- ii. What is the octet rule? How does ionic bonding happen?
- iii. Naming and writing formulas, including those with transition metals
 1. $\text{MgCl}_2 =$ _____
 2. $\text{KF} =$ _____
 3. Iron (III) oxide = _____
- iv. Common polyatomic ions
 1. $\text{Na}_2\text{CO}_3 =$ _____
 2. Aluminum hydroxide = _____

III. UNIT 5: COVALENT BONDING

- a. What is the octet rule? How does covalent bonding happen?
- b. Single, double, triple bonds + valence electrons
- c. Naming and writing formulas (system of prefixes)
- d. VSEPR: predicting shapes of molecules
- e. Polar/nonpolar: what is it, how do you know?
 - i. Draw Lewis structures for the following molecules and identify the shape and polarity:
 1. H_2O
 2. NH_3
 3. CH_4
 4. SF_6
 5. SO_4^{2-}

IV. UNIT 6/7: CHEMICAL REACTIONS

- a. What is a chemical reaction?

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Clifford Lane



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