

Name Tanya BenedictDate 5/19/24

Percent Yield Practice Problems

Directions: Determine the percent yield for each scenario in the problems below. You may need to determine the theoretical yield using stoichiometry before determining the percent yield.

$$\text{Percent Yield} = \frac{\text{Actual Yield}}{\text{Theoretical Yield}} \times 100$$

1. Determine the percent yield if a reaction yielded 23.5 grams of water vapor. The theoretical yield of the reaction was 32.0 grams of water vapor.

$$\frac{23.5 \text{ g}}{32 \text{ g}} \times 100 = 73.4\%$$

2. A student performs an experiment and generates 5.0 grams of hydrogen gas. They calculate the theoretical yield of their reaction and determine it should have generated 7.5 grams of hydrogen gas. What was the percent yield for the student's experiment?

$$\frac{5 \text{ g}}{7.5} \times 100 = 66.7\%$$

3. An industrial-grade reaction generates 4,500 kg of product. The company anticipated a yield of 5,500 kg. What is the percent yield for this reaction?

$$\frac{4500 \text{ kg}}{5500 \text{ kg}} \times 100 = 81.8\%$$

4. A pharmaceutical drug company has a prescription drug that has a percent yield of 76%. If the company uses enough materials to theoretically generate 8,000 kg of product, how much will they actually produce?

$$\frac{76}{100} \times 8000 \text{ kg} = 6080 \text{ kg}$$

Practice Percent Yield 44

Owen Mansfield, Sam Holyman

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author names.

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Calculating Theoretical Yield Is Not Impossible! Michael S. Leonard, 2025-02-25 Calculating theoretical and percent yield is a fundamental skill for the laboratory This book primarily targets Organic Chemistry Laboratory courses at the high school or college and university level as a supplemental resource to help students master this skill It begins with simple examples from everyday life demonstrates the importance of balancing the equation addresses the role of the mole in these computations discusses different types of liquids considers the role of significant figures and culminates with the planning of syntheses There are suggestions for further reading as well as practice problems and questions to ensure mastery Begins with examples from everyday life that enable students to understand the concepts of theoretical and percent yield before applying those concepts to the laboratory Addresses the necessity of balancing the reaction equation the centrality of the mole in these calculations and the role of significant figures

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