

# Periodic Trends Practice



## Atomic Radius

The distance from the nucleus of one bonded atom.



## Electronegativity

How much an atom "wants" electrons or the atom's tendency to attract electrons.



## Ionization Energy

The amount of energy required to remove electrons from an atom.



## Electron Affinity

The amount of energy released when a neutral gaseous atom gains an electron.



## Driving Question

How can the periodic table be used to predict atomic characteristics such as atomic radius, electronegativity, ionization energy, and electron affinity?

Describe what causes this trend...

As you move to the right, the number of protons increases, but the number of electrons also increases. This causes a stronger attraction between the nucleus and the electrons, pulling them closer.

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**Shielding?**

What is shielding? What effect does it have on the periodic trends defined above?

Shielding occurs when other electrons block out the positive pull of the nucleus, making it harder for the electrons. This causes a decrease in atomic radius, electronegativity, ionization energy, and electron affinity.

Add arrows to each of the following periodic tables to indicate the direction of the trend as it increases or decreases from left to right and top to bottom. (For Fig. 1 use a different color for each of the periodic tables so you can easily remember them.)



# Practice Identifying Periodic Trends

**Daniel F McAuley**



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