

STUDY GUIDE 11: Oxidation and Reduction (Chapter 12)

Remember: Oxidation = Loss of Electrons

Oxidation: loss of electrons; decrease in O.N. number (including 0, 0, 0 bonds)

Reduction: gain of electrons; increase in O.N. number; decrease in O.N. number



Be careful: some compounds contain both oxidizing and reducing agents, but in a single reaction.



REDOX REACTIONS

Redox reactions involve the transfer of electrons.

Example:

Consider the reaction between iron(II) sulfate and potassium dichromate in acidic solution.



Procedure: To find the oxidation state of each element in the reactants and products, use the following rules:



Oxidation And Reduction Study Guide

Richard Bailey



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