

Answer key-H.W(2)-Oxidation-reduction reaction

1-Multiple choice questions:

1-Oxidation involves

- a-gaining electrons; gaining oxygen
- b-losing electrons; losing oxygen
- c-gaining electrons; losing oxygen
- d-losing electrons; gaining oxygen

2- Reduction involves

- a-gaining electrons; gaining oxygen
- b-losing electrons; losing oxygen
- c-gaining electrons; losing oxygen
- d-losing electrons; gaining oxygen

3-In the reaction



- a-Reduced
- b- Synthesized
- c-Oxidized

4- In the reaction $\text{Zn} + \text{H}_2\text{O} \rightarrow \text{ZnO}_2 + \text{H}_2$

which element, is oxidized?

- a-Zinc
- b- Hydrogen
- c- Oxygen

5- Which substance is reduced in the following reaction?



- a-Li
- b-Na
- c-O
- d-H

6-Which of the following is the balanced half -reaction for the oxidation of Cu to Cu^{+2} ?

- a- $\text{Cu} \rightarrow \text{Cu}^{+2}$
- b- $\text{Cu} + 2 e^- \rightarrow \text{Cu}^{+2}$
- c- $\text{Cu} \rightarrow \text{Cu}^{+2} + 2 e^-$

7-Another name for a "oxidation-reduction" reaction is

- a-chemical reaction
- b-neutralization reaction
- c-redox reaction

Remember

Redox (Oxidation and Reduction)

Oxidation	Reduction
Gain oxygen ↑	Loss oxygen ↓
Gain oxidation Number ↑	Loss oxidation Number ↓
Loss hydrogen ↓	Gain hydrogen ↑
Loss electron ↓	Gain electron ↑

Oxidation And Reduction Packet Answer Key

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Principles of Oxidation and Reduction A. G. Sharpe,1973 **A Study of Reversible Oxidation and Reduction Reactions in Solutions** Eugene Paul Schoch,1904 **Oxidation and Reduction** ,1992 *Oxidation Reduction Reactions in Aqueous Solutions* Stanley Duckworth,1962 **Oxidation-reduction Reactions in Solution** Keith James Ellis,1969

Keys to Oxidation-reduction ,1974 **Oxidation-reduction Reactions in Inorganic Chemistry** Eric Randolph Jette,1927 **The Mechanisms of Oxidation-reduction Reactions in Solution** Norman Sutin,1967

Oxidation-reduction ,1975 **Oxidation and Reduction in Inorganic and Analytical Chemistry** Alan Vincent,1985-11-22 Beginning with basic principles this self instructional text leads students to an advanced understanding of oxidation and reduction Covers use of the mole concept and equation balancing and elementary thermodynamics concepts

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