

- According to the Bronsted—Lowry definition, which chemical species can function *both* as an acid and as a base?
  - $\text{Cl}^-$
  - $\text{SO}_4^{2-}$
  - $\text{NH}_4^+$
  - $\text{HCO}_3^-$
  - $\text{H}_3\text{O}^+$
- In the reaction  $\text{CN}^- + \text{H}_2\text{O} \rightleftharpoons \text{HCN} + \text{OH}^-$  which is an acid—base conjugate pair?
  - $\text{H}_2\text{O}$  and  $\text{HCN}$
  - $\text{H}_2\text{O}$  and  $\text{OH}^-$
  - $\text{CN}^-$  and  $\text{H}_2\text{O}$
  - $\text{HCN}$  and  $\text{OH}^-$
- Given that  $\text{HX}$  is a stronger Bronsted acid than  $\text{HY}$  in aqueous solution, which is true of a 1 M solution of  $\text{NaX}$ ?
  - It is less basic than a 1 M solution of  $\text{NaY}$ .
  - It is more basic than a 1 M solution of  $\text{NaY}$ .
  - It yields a neutral solution.
  - It is more concentrated than a 1 M solution of  $\text{NaY}$ .
- The pH of a 0.03 M  $\text{HCl}$  solution is
  - 1.5
  - 2.5
  - 3.5
  - 12.5
- The pH of a  $1.0 \times 10^{-2}$  M  $\text{Ba}(\text{OH})_2$  solution at  $25^\circ\text{C}$  is
  - 2.7
  - 3.0
  - 11.0
  - 11.3
- Which is the strongest acid?
  - $\text{HClO}$
  - $\text{HClO}_2$
  - $\text{HClO}_3$
  - $\text{HClO}_4$
- Which salt reacts with water (hydrolyzes) to produce a basic solution?
  - $\text{NaC}_2\text{H}_3\text{O}_2$
  - $\text{NaNO}_3$
  - $\text{NH}_4\text{Cl}$
  - $\text{BaSO}_4$

- The *weakest* of the bases listed is
 

Acid	Conjugate Base	$K_a$ (Ionization Constant of Acid)
$\text{HCl}$	$\text{Cl}^-$	100% ionized
$\text{HSO}_4^-$	$\text{SO}_4^{2-}$	$1.2 \times 10^{-2}$
$\text{H}_2\text{S}$	$\text{HS}^-$	$5.7 \times 10^{-8}$
$\text{HS}^-$	$\text{S}^{2-}$	$1.2 \times 10^{-13}$

  - $\text{Cl}^-$
  - $\text{CN}^-$
  - $\text{HS}^-$
  - $\text{S}^{2-}$
  - $\text{SO}_4^{2-}$

- The oxide of which element will react with water to form the strongest acid?
 

	Main Groups							
Period	I	II	III	IV	V	VI	VII	O
First								
Second	X	Y	Z	P		Q	S	U
Third	W					R	T	M

  - W
  - M
  - P
  - R
  - Z

- Which statement is a logical inference from the fact that a 0.10 M solution of potassium acetate  $\text{KC}_2\text{H}_3\text{O}_2$  is less alkaline than a 0.10 M solution of potassium cyanide,  $\text{KCN}$ ?
  - Hydrocyanic acid is a weaker acid than acetic acid.
  - Hydrocyanic acid is less soluble in water than acetic acid.
  - Cyanides are less soluble than acetates.
  - Acetic acid is a weaker acid than hydrocyanic acid.
- In the titration of 50.0 mL of 0.100 M benzoic acid (a monoprotic acid) with 50.0 mL of 0.100 M  $\text{NaOH}$ , the properties of the solution at the equivalence point will correspond exactly to the properties of
  - a 0.100 M sodium solution.
  - a 0.0500 M sodium hydroxide solution.
  - a 0.0500 M benzoic acid solution.
  - a 0.0500 M sodium benzoate solution.

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**Robert R. Redfield**



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