

11

CHEMICAL REACTIONS

Practice Problems

In your notebook, solve the following problems. Use the 3-step problem-solving approach you learned in Chapter 1.

SECTION 11.1 DESCRIBING CHEMICAL REACTIONS

1. Write the skeleton equation for the reaction between hydrogen and oxygen that produces water.
2. Write the skeleton equation for the reaction that produces iron(II) sulfide from iron and sulfur.
3. Write the skeleton equation representing the heating of magnesium carbonate to produce solid magnesium oxide and carbon dioxide gas.
4. Write a balanced equation for the production of HCl gas from its elements.
5. Write a sentence that completely describes the chemical reaction represented by this balanced equation.



6. Write the word equation for the following equation. Write a sentence fully describing the reaction. Is the equation correctly balanced? Explain.



7. Write a balanced equation representing the formation of aqueous sulfuric acid from water and sulfur trioxide gas.
8. Write a balanced equation from this word equation.
aqueous silver nitrate + copper metal $\rightarrow$ silver metal + aqueous copper nitrate
9. Write a balanced equation for the following word equation.



SECTION 11.2 TYPES OF CHEMICAL REACTIONS

1. Write a balanced equation representing the reaction of magnesium with oxygen gas to produce magnesium oxide.
2. Write the balanced equation for the reaction that occurs between aluminum and fluorine.
3. Write the balanced equation for the production of oxygen gas and potassium chloride from the decomposition of potassium chlorate.
4. Write the balanced equation for the reaction between hydrochloric acid and calcium metal. The products are hydrogen gas and calcium chloride.
5. Write the balanced equation for the combustion of propane (C_3H_8) to produce carbon dioxide and water vapor.
6. Write the balanced equation for the reaction between iron(III) chloride and sodium hydroxide. The products are iron(III) hydroxide and sodium chloride.

Practice Problems Describing Chemical Reactions

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Practice Problems Describing Chemical Reactions:

The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises **ATI TEAS 7 Crash Course with Online Practice Test, 4th Edition** John Allen, 2024-01-11

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Education John K. Gilbert, David Treagust, 2009-02-28 Chemistry seeks to provide qualitative and quantitative explanations for the observed behaviour of elements and their compounds Doing so involves making use of three types of representation the macro the empirical properties of substances the sub-micro the natures of the entities giving rise to those properties and the symbolic the number of entities involved in any changes that take place Although understanding this triplet relationship is a key aspect of chemical education there is considerable evidence that students find great difficulty in achieving mastery of the ideas involved In bringing together the work of leading chemistry educators who are researching the triplet relationship at the secondary and university levels the book discusses the learning involved the problems that students encounter and successful approaches to teaching Based on the reported research the editors argue for a coherent model for understanding the triplet relationship in chemical education Mathematical Models of Chemical Reactions Péter Érdi, János Tóth, 1989

NASA Tech Briefs, 1995

LSSENS, a General Chemical Kinetics and Sensitivity Analysis Code for Homogeneous

Gas-phase Reactions. 2: Code Description and Usage Krishnan Radhakrishnan,1994 **Chemical Kinetics in Combustion and Reactive Flows: Modeling Tools and Applications** V. I. Naoumov,V. G. Krioukov,A. L. Abdullin,A. V. Demin,2019-08-22 Introduces advanced mathematical tools for the modeling simulation and analysis of chemical non equilibrium phenomena in combustion and flows following a detailed explanation of the basics of thermodynamics and chemical kinetics of reactive mixtures Researchers practitioners lecturers and graduate students will find this work valuable
Concrete and Constructional Engineering ,1924

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