

Chemical Equilibrium

AP Chem Unit 13: Practice Problems

Practice Problem #1: Write the equilibrium expression for the following reaction:



Practice Problem #2: The following equilibrium concentrations were observed for the Haber process for the synthesis of ammonia at 127°C:

$$[\text{NH}_3] = 3.1 \times 10^{-2} \text{ mol/l}$$

$$[\text{N}_2] = 8.5 \times 10^{-1} \text{ mol/l}$$

$$[\text{H}_2] = 3.1 \times 10^{-3} \text{ mol/l}$$

a) Calculate the value of K at 127°C for this reaction.

b) Calculate the value of the equilibrium constant at 127°C for the reaction:



c) Calculate the value of the equilibrium constant at 127°C for the reaction:



Practice Problems Chemical Equilibrium

VM Jensen

A decorative graphic element consisting of a light blue horizontal bar with a rounded right end, and a red circular gradient shape partially visible behind it.

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