

Naming Ionic Compounds Practice Worksheet

Name the following ionic compounds:

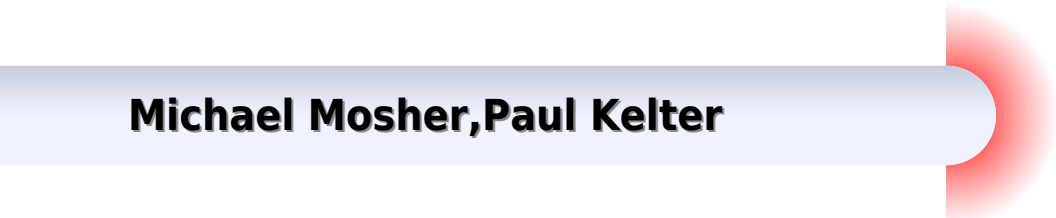
- 1) NH_4Cl Ammonium chloride
- 2) $\text{Fe}(\text{NO}_3)_3$ Iron (III) Nitrate
- 3) TiBr_3 Titanium (III) Bromide
- 4) Cu_3P Copper (I) Phosphide
- 5) SnSe_2 Tin (IV) Selenide
- 6) GaAs Gallium Arsenide
- 7) $\text{Pb}(\text{SO}_4)_2$ Lead (IV) Sulfate
- 8) $\text{Be}(\text{HCO}_3)_2$ Beryllium Hydrogen Carbonate
- 9) $\text{Mn}_2(\text{SO}_3)_3$ Manganese (II) Sulfite
- 10) $\text{Al}(\text{CN})_3$ Aluminum Cyanide

Write the formulas for the following compounds:

- 11) chromium (VI) phosphate $\text{Cr}(\text{PO}_4)_2$
- 12) vanadium (IV) carbonate $\text{V}(\text{CO}_3)_2$
- 13) tin (II) nitrite $\text{Sn}(\text{NO}_2)_2$
- 14) cobalt (III) oxide Co_2O_3
- 15) titanium (II) acetate $\text{Ti}(\text{C}_2\text{H}_3\text{O}_2)_2$
- 16) vanadium (V) sulfide V_2S_5
- 17) chromium (III) hydroxide $\text{Cr}(\text{OH})_3$
- 18) lithium iodide LiI
- 19) lead (II) nitride Pb_3N_2
- 20) silver bromide AgBr

Practice Ionic And Polyatomic Compounds Answers

Michael Mosher, Paul Kelter



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