

## Acids and Bases Worksheet

1. Acids and bases share some properties but are different in others. Use the samples at the side of the class to complete the chart. If both acids and bases share the property, place it in both columns.

| Property                         | Acids               | Bases    |
|----------------------------------|---------------------|----------|
| Conductivity in Aqueous Solution | very                | very     |
| Taste                            | Sour                | Bitter   |
| Feel                             | —                   | Slippery |
| Reaction with Metal              | produces $H_2$ gas  | —        |
| Reaction with Carbonate          | produces $CO_2$ gas | —        |
| Colour on Litmus Paper           | Red                 | Blue     |
| Colour with Universal Indicator  | Red                 | Purple   |
| Colour with phenolphthalein      | Colourless          | pink     |

2. Indicate if each compound is an acid or base and name it.

|            | Acid/Base? | Name                        |
|------------|------------|-----------------------------|
| $HNO_3$    | <u>A</u>   | <u>nitric acid</u>          |
| $Mg(OH)_2$ | <u>B</u>   | <u>magnesium hydroxide</u>  |
| $H_3PO_4$  | <u>A</u>   | <u>phosphoric acid</u>      |
| $H_2CO_3$  | <u>A</u>   | <u>carbonic acid</u>        |
| HF         | <u>A</u>   | <u>hydrofluoric acid</u>    |
| $Fe(OH)_3$ | <u>B</u>   | <u>iron (iii) hydroxide</u> |

3. Write the formula for each of the following acids/bases:

|                    |                              |                   |                              |
|--------------------|------------------------------|-------------------|------------------------------|
| hydrochloric acid  | <u>HCl</u>                   | hydrochloric acid | <u>HCl</u>                   |
| aluminum hydroxide | <u><math>Al(OH)_3</math></u> | sodium hydroxide  | <u>NaOH</u>                  |
| lithium hydroxide  | <u>LiOH</u>                  | calcium hydroxide | <u><math>Ca(OH)_2</math></u> |
| sulfuric acid      | <u><math>H_2SO_4</math></u>  | sulfurous acid    | <u><math>H_2SO_3</math></u>  |

# Practice Acid And Bases

**H Kauffman**



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