

Stoichiometry

Lecture Examples Answer Key

Ex. 1 Balance the following chemical equations:



Ex. 2 Write a balanced chemical reaction for the combustion of octane, $\text{C}_8\text{H}_{18(l)}$, in air.



Ex. 3 Write the balanced chemical equation for the combination of the metallic element calcium with the nonmetallic element oxygen, O_2 .



Ex. 4 Calculate the molecular weight of

(a) N_2O_5 **MW = 108.0104 g mol⁻¹**

(b) $\text{Ca}(\text{C}_2\text{H}_3\text{O}_2)_2$ **MW = 158.167 g mol⁻¹**

(c) $[\text{K}_2(\text{UO}_2)_2(\text{VO}_4)_2 \cdot 3\text{H}_2\text{O}]$ **MW = 902.1758 g mol⁻¹**

Ex. 5 Calculate the percentage by mass of oxygen in 4a and 4b above.

(a) $[(5 \times 15.9994 \text{ g mol}^{-1}) / 108.0104 \text{ g mol}^{-1}] \times 100\% = \mathbf{74.0642\%}$

(b) $[(4 \times 15.9994 \text{ g mol}^{-1}) / 158.167 \text{ g mol}^{-1}] \times 100\% = \mathbf{40.4620\%}$

Ex. 6 (a) Calculate the number of O atoms in 0.470 mol of $\text{C}_6\text{H}_{12}\text{O}_6$.

$$\begin{aligned} \# \text{ O atoms} &= (0.470 \text{ mol } \text{C}_6\text{H}_{12}\text{O}_6)(6.022 \times 10^{23} \text{ molecules/1 mol})(6 \text{ O atoms / 1 molecule}) \\ &= \mathbf{1.70 \times 10^{24} \text{ O atoms}} \end{aligned}$$

(b) Calculate the total number of ions in 38.1 g of CaF_2 .

$$\begin{aligned} \# \text{ ions} &= (38.1 \text{ g } \text{CaF}_2)(1 \text{ mol} / 78.075 \text{ g})(6.022 \times 10^{23} \text{ molecules/1 mol})(3 \text{ ions / 1 molecule}) \\ &= \mathbf{8.82 \times 10^{23} \text{ ions}} \end{aligned}$$

Ex. 7 (a) What is the mass in grams of 1 mol of glucose, $\text{C}_6\text{H}_{12}\text{O}_6$?

$$\text{mass} = (1 \text{ mol})(180.1572 \text{ g/1 mol}) = \mathbf{180.1572 \text{ g}}$$

(b) What is the mass in grams of 3.52 mol of chromium(III) sulfate decahydrate?

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Katrin Zwirgmaier



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CALCULATIONS K.V. NARAYANAN ,B. LAKSHMIKUTTY,2016-12-01 Designed as a textbook for the undergraduate students of chemical engineering and related disciplines such as biotechnology polymer technology petrochemical engineering electrochemical engineering environmental engineering and safety engineering the chief objective of the book is to prepare students to make analysis of chemical processes through calculations and to develop systematic problem solving skills in them The text presents the fundamentals of chemical engineering operations and processes in a simple style that helps the students to gain a thorough understanding of chemical process calculations The book deals with the principles of stoichiometry to formulate and solve material and energy balance problems in processes with and without chemical reactions With the help of examples the book explains the construction and use of reference substance plots equilibrium diagrams psychrometric charts steam tables and enthalpy composition diagrams It also elaborates on thermophysics and thermochemistry to acquaint the students with the thermodynamic principles of energy balance calculations The book is supplemented with Solutions Manual for instructors containing detailed solutions of all chapter end unsolved problems NEW TO THE SECOND EDITION Incorporates a new chapter on Bypass Recycle and Purge Operations Comprises updations in some sections and presents new sections on Future Avenues and Opportunities in Chemical Engineering Processes in Biological and Energy Systems Contains several new worked out examples in the chapter on Material Balance with Chemical Reaction Includes GATE questions with answers up to the year 2016 in Objective type questions KEY FEATURES SI units are used throughout the book All basic chemical engineering operations and processes are introduced and different types of problems are illustrated with worked out examples Stoichiometric principles are extended to solve problems related to bioprocessing environmental engineering etc Exercise problems more than 810 are organised according to the difficulty level and all are provided with answers **Jacaranda Chemistry 2 VCE Units 3 and 4, 3e learnON and Print** Robert Stokes,Angela Stubbs,Neale Taylor,Jason Bourke,Ben Williams,Maida Derbogolian,2023-12-26 Developed by expert Victorian teachers for VCE students The NEW Jacaranda Chemistry VCE series continues to deliver curriculum aligned material that caters to students of all abilities Our expert author team of practising teachers and assessors ensures 100% coverage of the

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Learning with Understanding in the Chemistry Classroom Iztok Devetak, Saša Aleksij Glažar, 2014-01-14 This volume offers a critical examination of a variety of conceptual approaches to teaching and learning chemistry in the school classroom Presenting up to date research and theory and featuring contributions by respected academics on several continents it explores ways of making knowledge meaningful and relevant to students as well as strategies for effectively communicating the core concepts essential for developing a robust understanding of the subject Structured in three sections the contents deal first with teaching and learning chemistry discussing general issues and pedagogical strategies using macro sub micro and symbolic representations of chemical concepts Researchers also describe new and productive teaching strategies The second section examines specific approaches that foster learning with understanding focusing on techniques such as cooperative learning presentations laboratory activities multimedia simulations and role playing in forensic chemistry classes The final part of the book details learner centered active chemistry learning methods active computer aided learning and trainee chemistry teachers use of student centered learning during their pre service education Comprehensive and highly relevant this new publication makes a significant contribution to the continuing task of making chemistry classes engaging and effective

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