

Quiz: Oxidation of Elements

Instructions: Choose the best answer for each question. Each question assesses your understanding of the oxidation of elements and related concepts.

1. What is oxidation in terms of electron transfer?

- A) Gain of electrons
- B) Loss of electrons
- C) Gain of protons
- D) Loss of neutrons

2. Which of the following reactions represents oxidation?

- A) $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
- B) $\text{Fe}^{2+} + 2\text{e}^- \rightarrow \text{Fe}$
- C) $2\text{Na} + \text{Cl}_2 \rightarrow 2\text{NaCl}$
- D) $\text{Cu} + 2\text{Ag}^+ \rightarrow \text{Cu}^{2+} + 2\text{Ag}$

3. In a redox reaction, what does the term "reducing agent" refer to?

- A) The substance that gets oxidized
- B) The substance that gets reduced
- C) The substance that remains unchanged
- D) The substance that produces energy

4. Which element is most likely to be oxidized in the reaction: $2\text{Fe} + \text{Cu}^{2+} \rightarrow 2\text{Fe}^{2+} + \text{Cu}$?

- A) Fe
- B) Cu^{2+}
- C) 2Fe^{2+}
- D) Cu

5. What is the oxidation state of sulfur in H_2SO_4 ?

- A) +2
- B) +4
- C) +6
- D) 0

6. Which of the following statements is true regarding oxidation states?

- A) Oxidation states can only be positive.
- B) The oxidation state of oxygen is always -2.
- C) Oxidation states are always whole numbers.

Oxidation State Word Answer Key

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