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web oxidation reduction redox reactions oxidation $\text{Cu} + 2\text{H}^+ \rightarrow \text{Cu}^{2+} + \text{H}_2$ reduction $\text{MnO}_4^- + 4\text{H}^+ + 3\text{e}^- \rightarrow \text{MnO}_2 + 2\text{H}_2\text{O}$ l when the half reactions listed above are combined into a balanced chemical equation the ratio of the coefficients $\text{Cu} + \text{MnO}_4^- + 4\text{H}^+ \rightarrow \text{Cu}^{2+} + \text{MnO}_2 + 2\text{H}_2\text{O}$ is

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web google classroom for the following redox equation $\text{MnO}_4^- + \text{H}^+ + \text{SO}_2 \rightarrow \text{Mn}^{2+} + \text{H}_2\text{O} + \text{SO}_4^{2-}$ what will be the stoichiometric coefficient of H^+ if the above equation is balanced in an acidic medium stuck use a hint report a problem do 4 problems

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web balance this redox reaction by the half reaction method by assuming an acidic solution $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$ balance the redox reaction in exercise 9 by the half reaction method by assuming a basic solution

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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^- + \text{Au} + \text{CN}^- + 2\text{e}^- + 2\text{H}^+ + \text{O}_2 \rightarrow 2\text{HCN} + 2\text{H}_2\text{O}_2$ make electrons equal convert to basic solution $4\text{CN}^- + 2\text{Au} + 2\text{CN}^- + 2\text{e}^-$ multiplied by a factor of 2 $2\text{e}^- + 2\text{H}_2\text{O} + \text{O}_2 \rightarrow 2\text{H}_2\text{O}_2 + 2\text{OH}^-$ the final answer

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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance c e o by adding c e h₂o

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web when balancing equations for redox reactions occurring in basic solution it is often necessary to add oh⁻ ions or the oh⁻ h₂o pair to fully balance the equation in this video we ll walk through this process for the reaction between clo⁻ and cr³⁺ in basic solution created by jay

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web bonus problem here is a balanced net ionic equation $3\text{Br}^- + 2\text{H}_2\text{O} + 5\text{Br}^- \rightarrow 3\text{H}_2\text{O} + 2\text{O}^{2-} + 8\text{Br}^-$ the above equation takes place in two stages first bromine reacts with OH^- ions to form bromide ions and BrO^- ions second the BrO^- ions react to form bromide ions and bromate BrO_3^- ions

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1 $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{Cr}^{3+} + \text{Fe}^{3+}$ solution 1 balanced half reactions $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$ 2 equalize the electrons $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$ multiplied by a factor of 6

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equations one for oxidation and one for reduction the equation is balanced by adjusting coefficients and adding h₂ o₂ h₂ and e⁻ in this order balance the atoms in the equation apart from o and h

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively c₆ h₈ o₆ 2h₂no₂ c₆ h₆ o₆ 2h₂ o₂no choose 1 answer c₆ h₈ o₆ and no₂ a c₆ h₈ o₆ and no₂ no₂ and c₆ h₈ o₆ b no₂ and c₆ h₈ o₆ h and no₂ c h and

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