

Question	Expected Answers	Marks	Additional Guidance
3 a	<p>measured pH &gt; 1 OR <math>[H^+] &lt; 0.1 \text{ (mol dm}^{-3}\text{)}</math> ✓</p> <p><math>[H^+] = 10^{-0.4}</math> ✓</p> <p><math>K_a = \frac{[H^+][CH_3CH_2COO^-]}{[CH_3CH_2COOH]}</math> OR <math>\frac{[H^+]^2}{[CH_3CH_2COOH]}</math> ✓</p> <p>Calculate <math>K_a</math> from <math>\frac{[H^+]^2}{0.100}</math> ✓</p>	4	<p>ALLOW <math>C_2H_5</math> throughout question</p> <p>ALLOW <math>[H^+] &lt; [CH_3CH_2COOH]</math> OR <math>[H^+] &lt; [HA]</math>  ALLOW measured pH is higher than expected  ALLOW measured pH is not as acidic as expected  ALLOW a quoted pH value or range &gt; 1 and &lt; 7  OR between 1 and 7</p> <p>ALLOW <math>[H^+] = \text{antilog } -\text{pH}</math> OR <math>[H^+] = \text{inverse log } -\text{pH}</math></p> <p>ALLOW <math>\frac{[H^+][A^-]}{[HA]}</math> OR <math>\frac{[H^+]^2}{[HA]}</math></p> <p>IF <math>K_a</math> is NOT given and <math>K_a = \frac{[H^+]^2}{0.100}</math> is shown, award mark for <math>K_a</math> also  (I.e. <math>K_a = \frac{[H^+]^2}{0.100}</math> is automatically awarded the last 2 marks)</p>
b	<p>Marks are for correctly calculated values.  Working shows how values have been derived.</p> <p><math>[H^+] = 10^{-13.46} = 3.47 \times 10^{-14} \text{ (mol dm}^{-3}\text{)}</math> ✓</p> <p><math>[OH^-] = \frac{1.0 \times 10^{-14}}{3.47 \times 10^{-14}} = 0.29 \text{ (mol dm}^{-3}\text{)}</math> ✓</p>	2	<p>ALLOW <math>3.467368505 \times 10^{-14}</math> and correct rounding to <math>3.5 \times 10^{-14}</math></p> <p>ALLOW 0.28840315 and correct rounding to 0.29,  i.e. ALLOW 0.288</p> <p>ALLOW alternative approach using pOH:</p> <p>pOH = <math>14 - 13.46 = 0.54</math> ✓  <math>[OH^-] = 10^{-0.54} = 0.29 \text{ (mol dm}^{-3}\text{)}</math> ✓</p> <p>Correct answer gets BOTH marks</p>

# Ocr F325 June 2010 Question Paper

**Shasha Hu**



**Ocr F325 June 2010 Question Paper:**

## **Ocr F325 June 2010 Question Paper** Book Review: Unveiling the Power of Words

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### **Table of Contents Ocr F325 June 2010 Question Paper**

1. Understanding the eBook Ocr F325 June 2010 Question Paper
  - The Rise of Digital Reading Ocr F325 June 2010 Question Paper
  - Advantages of eBooks Over Traditional Books
2. Identifying Ocr F325 June 2010 Question Paper
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ocr F325 June 2010 Question Paper
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr F325 June 2010 Question Paper
  - Personalized Recommendations
  - Ocr F325 June 2010 Question Paper User Reviews and Ratings
  - Ocr F325 June 2010 Question Paper and Bestseller Lists
5. Accessing Ocr F325 June 2010 Question Paper Free and Paid eBooks

- Ocr F325 June 2010 Question Paper Public Domain eBooks
- Ocr F325 June 2010 Question Paper eBook Subscription Services
- Ocr F325 June 2010 Question Paper Budget-Friendly Options
- 6. Navigating Ocr F325 June 2010 Question Paper eBook Formats
  - ePub, PDF, MOBI, and More
  - Ocr F325 June 2010 Question Paper Compatibility with Devices
  - Ocr F325 June 2010 Question Paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ocr F325 June 2010 Question Paper
  - Highlighting and Note-Taking Ocr F325 June 2010 Question Paper
  - Interactive Elements Ocr F325 June 2010 Question Paper
- 8. Staying Engaged with Ocr F325 June 2010 Question Paper
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ocr F325 June 2010 Question Paper
- 9. Balancing eBooks and Physical Books Ocr F325 June 2010 Question Paper
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ocr F325 June 2010 Question Paper
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Ocr F325 June 2010 Question Paper
  - Setting Reading Goals Ocr F325 June 2010 Question Paper
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ocr F325 June 2010 Question Paper
  - Fact-Checking eBook Content of Ocr F325 June 2010 Question Paper
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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