



Chapter 1

Temperature, Thermal Energy, and Heat

Temperature: a measure of how hot or cold an object is, based on the average kinetic energy of the particles that make up the object. It is related to the average kinetic energy of the particles that make up the object.

How _____

Explain the concept _____

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How _____

1. Temperature is a measure of the average kinetic energy of the particles that make up an object.

2. The higher the temperature, the more kinetic energy the particles have.

3. Temperature is a measure of the average kinetic energy of the particles that make up an object.

4. If two objects are at the same temperature, they have the same average kinetic energy of the particles that make up the objects.

5. The absolute zero of temperature is the temperature at which the particles have no kinetic energy.

6. A thermometer is used to measure the specific heat of a substance.

7. Temperature is a measure of the average kinetic energy of the particles that make up an object.

8. Heat is a form of energy.

9. The heat capacity of an object is the amount of heat energy it takes to raise the temperature of the object by one degree Celsius.

10. The amount of energy it takes to raise the temperature of 1 kg of a substance by 1°C is called the specific heat of the substance.

11. Water has a high specific heat capacity.

12. Materials with a high specific heat capacity can store a lot of energy without a large change in temperature.

Thermal energy: the energy that is transferred from one object to another by heat.

13. The amount of thermal energy that is transferred from one object to another is called the heat energy.

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