

Chapter 5 The Periodic Table

Section 5.2 The Modern Periodic Table

(pages 130–138)

This section explains the organization of the modern periodic table and discusses the general properties of metals, nonmetals, and metalloids.

Reading Strategy (page 130)

Previewing Before you read, complete the table by writing two questions about the periodic table on pages 132–133. As you read, write answers to your questions. For more information on this reading strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

Questions About the Periodic Table	
Question	Answer
Students might ask: What does atomic mass mean?	Atomic mass is a value that depends on the distribution of an element's isotopes in nature and the masses of those isotopes.
Why are two columns of elements placed below the main body of the table?	Two kinds of elements are placed below the table to make the table more compact.

Other possible questions: Why are there two numbering systems for the columns? Why is period 7 incomplete?

The Periodic Law (pages 131–133)

1. Is the following sentence true or false? In the modern periodic table, elements are arranged by increasing number of protons.
_____ **true** _____
2. Explain why the number of elements per period varies. _____ **The number of elements per period varies because the number of available orbitals increases from energy level to energy level.** _____
3. Properties of elements repeat in a predictable way when atomic numbers are used to arrange elements into groups. This pattern of repeating properties is called the _____ **periodic law** _____.

Atomic Mass (page 134)

4. Label the four types of information supplied for chlorine in the diagram.

a.	17
b.	Cl
c.	Chlorine
d.	35.453

- Atomic number**
- Element symbol**
- Element name**
- Atomic mass**

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Table of Contents Section 5 2 The Modern Periodic Table Reading Guide

1. Understanding the eBook Section 5 2 The Modern Periodic Table Reading Guide
 - The Rise of Digital Reading Section 5 2 The Modern Periodic Table Reading Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Section 5 2 The Modern Periodic Table Reading Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section 5 2 The Modern Periodic Table Reading Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section 5 2 The Modern Periodic Table Reading Guide
 - Personalized Recommendations
 - Section 5 2 The Modern Periodic Table Reading Guide User Reviews and Ratings
 - Section 5 2 The Modern Periodic Table Reading Guide and Bestseller Lists
5. Accessing Section 5 2 The Modern Periodic Table Reading Guide Free and Paid eBooks

- Section 5 2 The Modern Periodic Table Reading Guide Public Domain eBooks
- Section 5 2 The Modern Periodic Table Reading Guide eBook Subscription Services
- Section 5 2 The Modern Periodic Table Reading Guide Budget-Friendly Options
- 6. Navigating Section 5 2 The Modern Periodic Table Reading Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Section 5 2 The Modern Periodic Table Reading Guide Compatibility with Devices
 - Section 5 2 The Modern Periodic Table Reading Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section 5 2 The Modern Periodic Table Reading Guide
 - Highlighting and Note-Taking Section 5 2 The Modern Periodic Table Reading Guide
 - Interactive Elements Section 5 2 The Modern Periodic Table Reading Guide
- 8. Staying Engaged with Section 5 2 The Modern Periodic Table Reading Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section 5 2 The Modern Periodic Table Reading Guide
- 9. Balancing eBooks and Physical Books Section 5 2 The Modern Periodic Table Reading Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section 5 2 The Modern Periodic Table Reading Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Section 5 2 The Modern Periodic Table Reading Guide
 - Setting Reading Goals Section 5 2 The Modern Periodic Table Reading Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Section 5 2 The Modern Periodic Table Reading Guide
 - Fact-Checking eBook Content of Section 5 2 The Modern Periodic Table Reading Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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