

Name:

Period:

Stoichiometry Unit Study Guide

Vocabulary section

Directions: Use the following words from the word bank to fill in the blanks that best fits the sentence/definitions. Each word is only used **ONCE**.

Word bank			
Stoichiometry	Grams	Molar Mass	Mole ratio
Mole number	Scientific notation	Law of conservation of mass	Balanced chemical equation

1. The _____ states that in a chemical reaction, mass is **not** created or destroyed. It also means that if an element is present among the reactants it must almost be present in the product.
2. _____ is the mass of one mole of any atom/molecule and has a unit of **g/mol**.
3. _____ is the unit that measures mass only.
4. _____ is the relationship between the amounts/quantities of substances taking part in a chemical reaction. We use the "train track method" to solve these types of problems.
5. The _____ is **1 mole = 6.02×10^{23} atoms/molecules**.
6. A _____ is like the **recipe** of a chemical reaction and must follow the law of conservation.
7. _____ is a way of writing very large or very small numbers to use them more effectively.
8. _____ is a **conversion factor** that relates the amounts in moles of any two substances in a chemical reaction.

Directions: Answer the following questions using dimensional analysis (AKA the train track method). Show all your work and include units!

1. Find the number of moles in 46.32 grams of H_2O .

Section Study Guide Stoichiometry Answers

AW Chickering



Section Study Guide Stoichiometry Answers:

The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

Wastewater Operator Certification Study Guide John Giorgi, 2011-01-12

Study Guide/Selected Solutions Manual Julie R. Frentrop, 2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book

Study Guide to Accompany Basics for Chemistry Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation

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